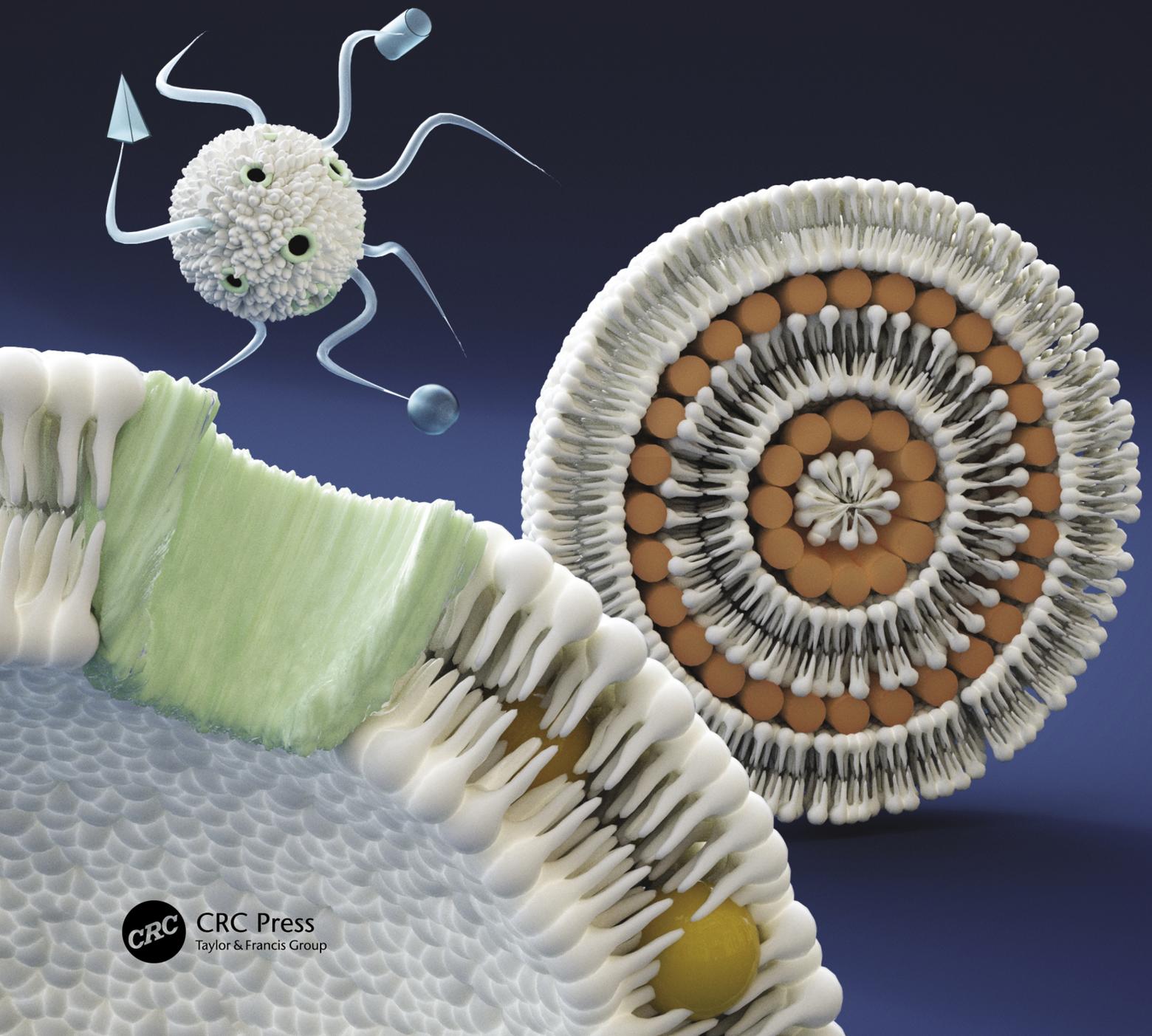


edited by Cyrus R. Safinya & Joachim O. Rädler

# Handbook of Lipid Membranes

Molecular, Functional, and Materials Aspects



# Handbook of Lipid Membranes

## Molecular, Functional, and Materials Aspects

Edited by  
Cyrus R. Safinya  
Joachim O. Rädler



**CRC Press**

Taylor & Francis Group  
Boca Raton London New York

---

CRC Press is an imprint of the  
Taylor & Francis Group, an **informa** business

First Edition published 2021  
by CRC Press  
6000 Broken Sound Parkway NW, Suite 300, Boca Raton, FL 33487-2742

and by CRC Press  
2 Park Square, Milton Park, Abingdon, Oxon, OX14 4RN

© 2021 Taylor & Francis Group, LLC

CRC Press is an imprint of Taylor & Francis Group, LLC

Reasonable efforts have been made to publish reliable data and information, but the author and publisher cannot assume responsibility for the validity of all materials or the consequences of their use. The authors and publishers have attempted to trace the copyright holders of all material reproduced in this publication and apologize to copyright holders if permission to publish in this form has not been obtained. If any copyright material has not been acknowledged, please write and let us know so we may rectify in any future reprint.

Except as permitted under U.S. Copyright Law, no part of this book may be reprinted, reproduced, transmitted, or utilized in any form by any electronic, mechanical, or other means, now known or hereafter invented, including photocopying, microfilming, and recording, or in any information storage or retrieval system, without written permission from the publishers.

For permission to photocopy or use material electronically from this work, access [www.copyright.com](http://www.copyright.com) or contact the Copyright Clearance Center, Inc. (CCC), 222 Rosewood Drive, Danvers, MA 01923, 978-750-8400. For works that are not available on CCC, please contact [mpkbookpermissions@tandf.co.uk](mailto:mpkbookpermissions@tandf.co.uk)

*Trademark notice:* Product or corporate names may be trademarks or registered trademarks and are used only for identification and explanation without intent to infringe.

ISBN: 978-1-466-55572-3 (hbk)  
ISBN: 978-1-032-01441-8 (pbk)  
ISBN: 978-0-429-19407-8 (ebk)

Typeset in Times  
by codeMantra

---

# 6 Charged Membranes: Poisson–Boltzmann Theory, The DLVO Paradigm, and Beyond

*Tomer Markovich and David Andelman*  
Tel Aviv University, Tel Aviv, Israel

*Rudolf Podgornik*  
University of Chinese Academy of Sciences, Beijing, China  
Institute of Physics, Chinese Academy of Sciences, Beijing, China  
Wenzhou Institute of the University of Chinese Academy of Sciences, Wenzhou, China

## CONTENTS

6.1	Introduction .....	100
6.2	Poisson–Boltzmann Theory .....	101
6.2.1	Debye–Hückel Approximation .....	103
6.3	One Planar Membrane .....	104
6.3.1	Counterions Only .....	104
6.3.2	Added Electrolyte .....	105
6.3.3	The Grahame Equation .....	107
6.4	Modified Poisson–Boltzmann Theory .....	107
6.5	Two-Membrane System: Osmotic Pressure .....	109
6.6	Two Symmetric Membranes, $\sigma_1 = \sigma_2$ .....	111
6.6.1	Counterions Only .....	111
6.6.2	Added Electrolyte .....	112
6.6.3	Debye–Hückel Regime .....	114
6.6.4	Intermediate Regime .....	114
6.6.5	Other Pressure Regimes .....	115
6.7	Two Asymmetric Membranes, $\sigma_1 \neq \sigma_2$ .....	115
6.7.1	The Debye–Hückel Regime .....	116
6.7.2	The Debye–Hückel Regime with Constant Surface Potential .....	116
6.7.3	Counterions Only .....	117
6.7.4	Attraction/Repulsion Crossover .....	118
6.8	Charge Regulation .....	118
6.8.1	Charge Regulation via Free Energy .....	121
6.9	van der Waals’ Interactions .....	122
6.9.1	The Hamaker Pairwise Summation .....	122
6.9.2	Macroscopic Theory of vdW Interactions .....	123
6.9.3	The Derjaguin–Landau–Verwey–Overbeek Theory .....	124
6.10	Limitations and Generalizations .....	125
	Notes .....	126
	Bibliography .....	126

## 6.1 INTRODUCTION

It is of great importance to understand electrostatic interactions and their key role in soft and biological matter. These systems typically consist of aqueous environment in which charges tend to dissociate and affect a wide variety of functional, structural, and dynamical properties. Among the numerous effects of electrostatic interactions, it is instructive to mention their effect on elasticity of flexible charged polymers (polyelectrolytes) and cell membranes, formation of self-assembled charged micelles, and stabilization of charged colloidal suspensions that results from the competition between repulsive electrostatic interactions and attractive van der Waals' interactions (Verwey and Overbeek 1948, Andelman 1995, 2005, Holm, Kekicheff and Podgornik 2000, Dean et al. 2014, Churaev, Derjaguin and Muller 2014).

In this chapter, we focus on charged membranes. Biological membranes are complex heterogeneous two-dimensional interfaces separating the living cell from its extracellular surrounding. Membranes also surround intercellular organelles such as the cell nucleus, Golgi apparatus, mitochondria, endoplasmic reticulum, and ribosomes. Electrostatic interactions control many of the membrane structural properties and functions, *e.g.*, rigidity, structural stability, lateral phase transitions, and their dynamics. Moreover, electric charges are a key player in processes involving more than one membrane such as membrane adhesion and cell–cell interaction, as well as the overall interactions of membranes with other intra- and extracellular proteins, biopolymers, and DNA.

How do membranes interact with their surrounding ionic solution? Charged membranes attract a cloud of oppositely charged mobile ions that forms a *diffuse electric double layer* (Gouy 1910, 1917, Chapman 1913, Debye and Hückel 1923, Verwey and Overbeek 1948, Israelachvili 2011). The system favours *local electroneutrality*, but while achieving it, entropy is lost. The competition between electrostatic interactions and entropy of ions in solution determines the exact distribution of mobile ions close to charged membranes. This last point shows the significance of temperature in determining the equilibrium properties, because temperature controls the strength of entropic effects as compared to electrostatic interactions. For soft materials, the thermal energy  $k_B T$  where  $k_B$  is the Boltzmann constant and  $T$  is the temperature, is also comparable with other characteristic energy scales associated with elastic deformations and structural degrees of freedom.

It is convenient to introduce a length scale for which the thermal energy is equal to the Coulombic energy between two unit charges. This is called the *Bjerrum length*, defined as follows:

$$\ell_B = \frac{e^2}{4\pi\epsilon_0\epsilon_w k_B T}, \quad (6.1)$$

where  $e$  is the elementary charge,  $\epsilon_0 = 8.85 \cdot 10^{-12} [\text{F/m}]$  is the vacuum permittivity,<sup>1</sup> and the dimensionless dielectric

constant of water is  $\epsilon_w = 80$  (for a complete list of symbols, see Table 6.1). The Bjerrum length is equal to about 0.7 nm at room temperatures,  $T = 300 \text{ K}$ .

A related length is the *Gouy–Chapman (GC) length* defined as

$$\ell_{GC} = \frac{2\epsilon_0\epsilon_w k_B T}{e|\sigma|} = \frac{e}{2\pi\ell_B|\sigma|} \sim \sigma^{-1}. \quad (6.2)$$

At this length scale, the thermal energy  $k_B T$  is equal to the Coulombic energy between a unit charge and a planar surface with a constant surface-charge density,  $\sigma$ . The Gouy–Chapman length  $\ell_{GC}$  is inversely proportional to  $\sigma$ . For strongly charged membranes,  $\ell_{GC}$  is rather small, on the order of a tenth of nanometre.

In their pioneering work of almost a century ago, Debye and Hückel introduced the important concept of screening of the electrostatic interactions between two charges in presence of all other cations and anions of the solution (Debye and Hückel 1923). This effectively limits the range of electrostatic interactions as will be further discussed below. The characteristic length for which the electrostatic interactions are screened is called the *Debye length*,  $\lambda_D$ , defined for monovalent 1:1 electrolyte, as

$$\lambda_D = \kappa_D^{-1} = (8\pi\ell_B n_b)^{-1/2} \approx \frac{0.3[\text{nm}]}{\sqrt{n_b[\text{M}]}}, \quad (6.3)$$

with  $n_b$  being the salt concentration (in molar) and  $\kappa_D$  being the inverse Debye length. The Debye screening length for 1:1 monovalent salts varies from about 0.3 nm in strong ionic solutions of 1 M to about 1  $\mu\text{m}$  in pure water, where the concentration of the dissociated  $\text{OH}^-$  and  $\text{H}^+$  ions is  $10^{-7} \text{ M}$ .

The aim of this chapter is to review some of the basic considerations underlying the behaviour of charged membranes in aqueous solutions using the three important length scales introduced above. We will not account for the detailed structure of real biological membranes, which can add considerable complexity, but restrict ourselves to simple model systems, relying on several assumptions and simplifications. The membrane is treated as a flat interface with a continuum surface charge distribution or constant surface potential. The mobile charge distributions are continuous, and we disregard the discreteness of surface charges that can lead to multipolar charge distributions.

This chapter is focussed only on static properties in thermodynamic equilibrium, excluding the interesting phenomena of dynamical fluctuations and dynamical responses to external fields (such as in electrochemistry systems). We mainly treat the mean-field approximation of the electric double-layer problem and the solutions of the classical Poisson–Boltzmann (PB) equation. Nevertheless, some effects of fluctuations and correlations will be briefly discussed in Section 6.10. We will also discuss the ionic finite size in Section 6.4, where the ‘modified PB equation’ is introduced.

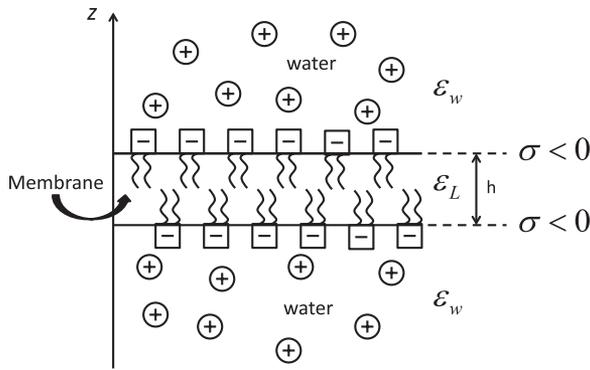
**TABLE 6.1**  
**Table of Symbols**

Symbol	Interpretation
$a$	Microscopic molecular size
$\alpha = \ln(a^3 K_d)$	Surface interaction parameter
$\alpha_i$	Polarizability of the $i^{\text{th}}$ molecule
$\beta = 1/k_B T$	Inverse thermal energy
$C_{\text{PB}}$	Differential capacitance of the PB model
$C_{\text{mPB}}$	Differential capacitance of the mPB model
$D$	Intermembrane separation
$\epsilon_w$	Water dimensionless dielectric constant
$\epsilon_L$	Lipid dimensionless dielectric constant
$F$	Helmholtz free energy
$\mathcal{F} = F(d) - F(d \rightarrow \infty)$	Excess Helmholtz free energy
$\phi_b = a^3 n_b$	Bulk volume fraction for monovalent electrolyte
$\phi_s =  \sigma  a^2 / e$	Surface area fraction
$h$	Membrane thickness
$\mathcal{H}$	Hamaker constant
$I = \frac{1}{2} \sum_{i=1}^M z_i^2 n_i^{(b)}$	Ionic strength
$K_d$	Kinetic constant
$\ell_B = e^2 / (4\pi\epsilon_0\epsilon_w k_B T)$	Bjerrum length
$\ell_{\text{GC}} = e / (2\pi\ell_B  \sigma )$ ; $\ell_{1,2} = e / (2\pi\ell_B  \sigma_{1,2} )$	Gouy–Chapman length
$\lambda_D = \kappa_D^{-1} = (8\pi\ell_B n_b)^{-1/2}$	Debye length
$\mu_i$	Intrinsic chemical potential
$\mu_i^{\text{tot}}$	Total chemical potential
$n_i(\mathbf{r})$	Concentration of the $i^{\text{th}}$ ionic species
$n_i^{(b)}$	Bulk concentration of the $i^{\text{th}}$ ionic species
$n_b$	Bulk concentration of monovalent electrolyte
$n_0$	Reference density, taken at zero potential
$n_i^{(s)}$ ; $n_s$	Surface density
$n_i^{(m)}$ ; $n_m$	Midplane density (two membranes)
$P$	Pressure
$\Pi = P_{\text{in}} - P_{\text{out}}$	Osmotic pressure
$\psi(\mathbf{r})$	Electrostatic potential
$\Psi(\mathbf{r}) = \beta e \psi(\mathbf{r})$	Dimensionless electrostatic potential
$\Psi_s$ ; $\psi_s$	Surface potential
$\Psi_m$ ; $\psi_m$	Midplane potential
$q_i = z_i e$	Charge of the $i^{\text{th}}$ ionic species
$\rho(\mathbf{r})$	Charge density of mobile ions
$\rho_f(\mathbf{r})$	Charge density of fixed charges
$\rho_{\text{tot}}(\mathbf{r})$	Total charge density
$\sigma$ ; $\sigma_{1,2}$	Surface charge density
$V$	Volume
$z_i$	Valency of the $i^{\text{th}}$ ionic species

The classical reference for the electric double layer is the book of Verwey and Overbeek (1948), which explains the DLVO (Derjaguin–Landau–Verwey–Overbeek) theory for stabilization of charged colloidal systems. More recent treatments can be found in many textbooks and monographs on colloidal science and interfacial phenomena, such as Evans and Wennerström (1999), Israelachvili (2011), and in two reviews by one of the present authors, Andelman (1995, 2005).

## 6.2 POISSON–BOLTZMANN THEORY

In Figure 6.1, a schematic view of a charged amphiphilic (phospholipid) membrane is presented. A membrane of thickness  $h \approx 4$  nm is composed of two monomolecular leaflets packed in a back-to-back configuration. The constituting molecules are amphiphiles having a charged ‘head’ and a hydrocarbon hydrophobic ‘tail’. For phospholipids, the



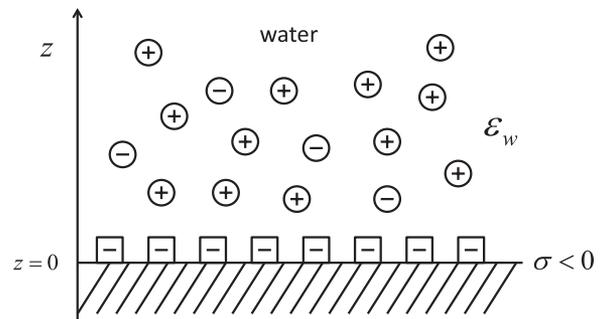
**FIGURE 6.1** A bilayer membrane of thickness  $h$  composed of two monolayers (leaflets), each having a negative charge density,  $\sigma < 0$ . The core membrane region (hydrocarbon tails) is modelled as a continuum medium with a dielectric constant  $\epsilon_L$ , while the embedding medium (top and bottom) is water and has a dielectric constant,  $\epsilon_w$ .

amphiphiles have a double tail. We model the membrane as a medium of thickness  $h$  having a dielectric constant,  $\epsilon_L$ , coming essentially from the closely packed hydrocarbon ('oily') tails. The molecular heads contribute to the surface charges, and the entire membrane is immersed in an aqueous solution characterized by another dielectric constant,  $\epsilon_w$ , assumed to be the water dielectric constant throughout the fluid.

The membrane charge can have two origins: either a charge group (*e.g.*,  $\text{H}^+$ ) dissociates from the polar head group into the aqueous solution, leaving behind an oppositely charged group in the membrane; or, an ion from the solution (*e.g.*,  $\text{Na}^+$ ) binds to a neutral site on the membrane and charges it (Borkovec, Jönsson, and Koper 2001). These association/dissociation processes are highly sensitive to the ionic strength and pH of the aqueous solution.

When the ionic association/dissociation is slow as compared with experimental times, the charges on the membrane can be considered as fixed and time independent, while for rapid association/dissociation, the surface charge can vary and is determined self-consistently from the thermodynamical equilibrium equations. We will further discuss the two processes of association/dissociation in Section 6.8. In many situations, the finite thickness of the membrane can be safely taken to be zero, with the membrane modelled as a planar surface as displayed in Figure 6.2. We will see later under what conditions this simplifying limit is valid.

Let us consider such an ideal membrane represented by a sharp boundary (located at  $z = 0$ ) that limits the ionic solution to the positive half space. The ionic solution contains, in general, the two species of mobile ions (anions and cations), and is modelled as a continuum dielectric medium as explained above. Thus, the boundary at  $z = 0$  marks the discontinuous jump of the dielectric constant between the ionic solution ( $\epsilon_w$ ) and the membrane ( $\epsilon_L$ ), which the ions cannot penetrate.



**FIGURE 6.2** Schematic illustration of a charged membrane, located at  $z=0$ , with charge density  $\sigma$ . Without loss of generality, we take  $\sigma < 0$ . For the counterion-only case, the surface charge is neutralized by the positive counterions. When monovalent (1:1) electrolyte is added to the reservoir, its bulk ionic density is  $n_{\pm}^{(b)} = n_b$ .

The PB equation can be obtained using two different approaches. The first is the one we present below combining the Poisson equation with the Boltzmann distribution, while the second one (presented later) is done through a minimization of the system free-energy functional. The PB equation is a mean-field (MF) equation, which can be derived from a field theoretical approach as the zeroth order in a systematic expansion of the grand-partition function (Podgornik and Žekš 1988, Borukhov, Andelman and Orland 1998, 2000, Netz and Orland 2000, Markovich, Andelman and Podgornik 2014, 2015, 2016, Markovich, Andelman and Orland 2016).

Consider  $M$  ionic species, each of them with charge  $q_i$ , where  $q_i = ez_i$  and  $z_i$  is the valency of the  $i^{\text{th}}$  ionic species. It is negative ( $z_i < 0$ ) for anions and positive ( $z_i > 0$ ) for cations. The mobile charge density (per unit volume) is defined as  $\rho(\mathbf{r}) = \sum_{i=1}^M q_i n_i(\mathbf{r})$  with  $n_i$  being the number density (per unit volume), and both  $\rho$  and  $n_i$  are continuous functions of  $\mathbf{r}$ .

In the MF approximation, each of the ions sees a local environment constituting of all other ions, which dictates a local electrostatic potential  $\psi(\mathbf{r})$ . The potential  $\psi(\mathbf{r})$  is a continuous function that depends on the total charge density through the Poisson equation:

$$\nabla^2 \psi(\mathbf{r}) = -\frac{\rho_{\text{tot}}(\mathbf{r})}{\epsilon_0 \epsilon_w} = -\frac{1}{\epsilon_0 \epsilon_w} \left[ \sum_{i=1}^M q_i n_i(\mathbf{r}) + \rho_f(\mathbf{r}) \right], \quad (6.4)$$

where  $\rho_{\text{tot}} = \rho + \rho_f$  is the total charge density and  $\rho_f(\mathbf{r})$  is a fixed external charge contribution.

As stated above, the aqueous solution (water) is modelled as a continuum featureless medium. This by itself represents an approximation because the ions themselves can change the local dielectric response of the medium (Ben-Yaakov, Andelman and Podgornik 2011, Levy, Andelman and Orland 2012, Adar et al. 2018) by inducing strong localized electric field. However, we will not include such refined local effects in this chapter.

The ions dispersed in solution are mobile and are allowed to adjust their positions. As each ionic species is in thermodynamic equilibrium, its density obeys the Boltzmann distribution:

$$n_i(\mathbf{r}) = n_i^{(b)} e^{-\beta q_i \psi(\mathbf{r})}, \quad (6.5)$$

where  $\beta = 1/k_B T$ , and  $n_i^{(b)}$  is the bulk density of  $i$ th species taken at zero reference potential,  $\psi = 0$ .

### BOLTZMANN DISTRIBUTION VIA ELECTROCHEMICAL POTENTIAL

A simple derivation of the Boltzmann distribution is obtained through the requirement that the *electrochemical potential* (total chemical potential)  $\mu_i^{\text{tot}}$ , for each ionic species, is constant throughout the system:

$$\mu_i^{\text{tot}} = \mu_i(\mathbf{r}) + q_i \psi(\mathbf{r}) = \text{const}, \quad (6.6)$$

where  $\mu_i(\mathbf{r})$  is the intrinsic chemical potential. For dilute ionic solutions, the  $i^{\text{th}}$  ionic species entropy is taken as an ideal gas one,  $\mu_i(\mathbf{r}) = k_B T \ln[n_i(\mathbf{r}) a^3]$ . By substituting  $a^3 n_i^{(b)} = \exp(\beta \mu_i^{\text{tot}})$  into Eq. (6.6), the Boltzmann distribution of Eq. (6.5) follows. This relation between the bulk ionic density and chemical potential is obtained by setting  $\psi = 0$  in the bulk and shows that one can consider the chemical potential,  $\mu_i^{\text{tot}}$ , as a Lagrange multiplier setting the bulk densities to be  $n_i^{(b)}$ . Note that we have introduced a microscopic length scale,  $a$ , defining a reference close packing density,  $1/a^3$ . Equation (6.6) assumes that the ions are point-like and have no other interactions in addition to their electrostatic one.

We now substitute Eq. (6.5) into Eq. (6.4) to obtain the *Poisson–Boltzmann equation*:

$$\nabla^2 \psi(\mathbf{r}) = -\frac{1}{\epsilon_0 \epsilon_w} \left[ \sum_{i=1}^M q_i n_i^{(b)} e^{-\beta q_i \psi(\mathbf{r})} + \rho_f(\mathbf{r}) \right]. \quad (6.7)$$

For binary monovalent electrolytes (denoted as 1:1 electrolyte),  $z_i = \pm 1$ , the PB equation reads

$$\nabla^2 \psi(\mathbf{r}) = \frac{1}{\epsilon_0 \epsilon_w} \left[ 2en_b \sinh(\beta e \psi(\mathbf{r})) - \rho_f(\mathbf{r}) \right]. \quad (6.8)$$

Generally speaking, the PB theory is a very useful analytical approximation with many applications. It is a good approximation at physiological conditions (electrolyte strength of about 0.1 M) and for other dilute monovalent electrolytes and moderate surface potentials and surface charge. Although the PB theory produces good results in these situations, it misses some important features associated with charge correlations and fluctuations of multivalent ions.

Moreover, close to a charged membrane, the finite size of the surface ionic groups and that of the counterions lead to deviations from the PB results (see Sections 6.4 and 6.8 for further details).

As the PB equation is a nonlinear equation, it can be solved analytically only for a limited number of simple boundary conditions. On the other hand, by solving it numerically or within further approximations or limits, one can obtain ionic profiles and free energies of complex structures. For example, the free energy change for a charged globular protein that binds onto an oppositely charged lipid membrane.

In an alternative approach, the PB equation can also be obtained by a minimization of the system free-energy functional. One can assume that the internal energy,  $U_{\text{el}}$ , is purely electrostatic and that the Helmholtz free energy,  $F = U_{\text{el}} - TS$ , is composed of an internal energy and an ideal mixing entropy,  $S$ , of a dilute solution of mobile ions.

The electrostatic energy,  $U_{\text{el}}$ , is expressed in terms of the potential  $\psi(\mathbf{r})$ :

$$\begin{aligned} U_{\text{el}} &= \frac{\epsilon_0 \epsilon_w}{2} \int_V d^3 r |\nabla \psi(\mathbf{r})|^2 \\ &= \frac{1}{2} \int_V d^3 r \left[ \sum_{i=1}^M q_i n_i(\mathbf{r}) \psi(\mathbf{r}) + \rho_f(\mathbf{r}) \psi(\mathbf{r}) \right], \end{aligned} \quad (6.9)$$

while the mixing entropy of ions is written in the dilute solution limit as

$$S = -k_B \sum_{i=1}^M \int_V d^3 r \left( n_i(\mathbf{r}) \ln[n_i(\mathbf{r}) a^3] - n_i(\mathbf{r}) \right). \quad (6.10)$$

Using Eqs. (6.9) and (6.10), the Helmholtz free energy can be written as

$$\begin{aligned} F &= \int_V d^3 r \left[ -\frac{\epsilon_0 \epsilon_w}{2} |\nabla \psi(\mathbf{r})|^2 + \left( \sum_{i=1}^M q_i n_i(\mathbf{r}) + \rho_f(\mathbf{r}) \right) \psi(\mathbf{r}) \right. \\ &\quad \left. + k_B T \sum_{i=1}^M \left( n_i(\mathbf{r}) \ln[n_i(\mathbf{r}) a^3] - n_i(\mathbf{r}) \right) \right], \end{aligned} \quad (6.11)$$

where the sum of the first two terms is equal to  $U_{\text{el}}$  and the third one is  $-TS$ . The variation of this free energy with respect to  $\psi(\mathbf{r})$ ,  $\delta F / \delta \psi = 0$ , gives the Poisson equation, Eq. (6.4), while from the variation with respect to  $n_i(\mathbf{r})$ ,  $\delta F / \delta n_i = \mu_i^{\text{tot}}$ , we obtain the electrochemical potential of Eq. (6.6). As before, substituting the Boltzmann distribution obtained from Eq. (6.6), into the Poisson equation, Eq. (6.4), gives the PB equation, Eq. (6.7).

### 6.2.1 DEBYE–HÜCKEL APPROXIMATION

A useful and quite tractable approximation to the nonlinear PB equation is its linearized version. For electrostatic potentials smaller than 25 mV at room temperature

(or equivalently  $e|\psi| < k_B T$ ,  $T \approx 300\text{K}$ ), this approximation can be justified and the well-known Debye–Hückel (DH) theory (Debye and Hückel 1923) is recovered. Linearization of Eq. (6.7) is obtained by expanding its right-hand side to the first order in  $\psi$ ,

$$\nabla^2 \psi(\mathbf{r}) = -\frac{1}{\epsilon_0 \epsilon_w} \sum_{i=1}^M q_i n_i^{(b)} + 8\pi \ell_B I \psi(\mathbf{r}) - \frac{1}{\epsilon_0 \epsilon_w} \rho_f(\mathbf{r}), \quad (6.12)$$

where  $I = \frac{1}{2} \sum_{i=1}^M z_i^2 n_i^{(b)}$  is the *ionic strength* of the solution.

The first term on the right-hand side of Eq. (6.12) vanishes because of electroneutrality of the bulk reservoir

$$\sum_{i=1}^M q_i n_i^{(b)} = 0, \quad (6.13)$$

recovering the DH equation:

$$\nabla^2 \psi(\mathbf{r}) = \kappa_D^2 \psi(\mathbf{r}) - \frac{1}{\epsilon_0 \epsilon_w} \rho_f(\mathbf{r}), \quad (6.14)$$

with the inverse Debye length,  $\kappa_D$ , defined as

$$\kappa_D^2 = \lambda_D^{-2} = 8\pi \ell_B I = 4\pi \ell_B \sum_{i=1}^M z_i^2 n_i^{(b)}. \quad (6.15)$$

For monovalent electrolytes,  $z_i = \pm 1$ ,  $\kappa_D^2 = 8\pi \ell_B n_b$  with  $n_i^{(b)} = n_b$ , and Eq. (6.3) is recovered. Note that the Debye length,  $\lambda_D = \kappa_D^{-1} \sim n_b^{-1/2}$ , is a decreasing function of the salt concentration.

The DH treatment gives a simple tractable description of the pair interactions between ions. It is related to the Green function associated with the electrostatic potential around a point-like ion that can be calculated by using Eq. (6.14) for a point-like charge,  $q$ , placed at the origin,  $\mathbf{r} = 0$ ,  $\rho_f(\mathbf{r}) = q\delta(\mathbf{r})$ ,

$$(\nabla^2 - \kappa_D^2)\psi(\mathbf{r}) = -\frac{q}{\epsilon_0 \epsilon_w} \delta(\mathbf{r}), \quad (6.16)$$

where  $\delta(\mathbf{r})$  is the Dirac  $\delta$ -function. The solution to the above equation can be written in spherical coordinates as

$$\psi(r) = \frac{q}{4\pi \epsilon_0 \epsilon_w r} e^{-\kappa_D r}. \quad (6.17)$$

It manifests the exponential decay of the electrostatic potential with a characteristic length scale,  $\lambda_D = 1/\kappa_D$ . In a crude approximation, this exponential decay is replaced by a Coulombic interaction, which is only slightly screened for  $r \leq \lambda_D$  and, thus, varies as  $\sim r^{-1}$ , while for  $r > \lambda_D$ ,  $\psi(r)$  is strongly screened and can sometimes be completely neglected.

### 6.3 ONE PLANAR MEMBRANE

We consider the PB equation for a single membrane, assumed to be planar and charged, and discuss separately two cases: (i) a charged membrane in contact with a solution containing only counterions, and (ii) a membrane in contact with a monovalent electrolyte reservoir.

As the membrane is taken to have an infinite extent in the lateral ( $x, y$ ) directions, the PB equation is reduced to an effective one-dimensional equation, where all local quantities, such as the electrostatic potential,  $\psi(\mathbf{r}) = \psi(z)$ , and ionic densities,  $n(\mathbf{r}) = n(z)$ , depend only on the  $z$ -coordinate perpendicular to the planar membrane.

For a binary monovalent electrolyte (1:1 electrolyte,  $z_i = \pm 1$ ), the PB equation from Eq. (6.7), reduces in its effective one-dimensional form to an ordinary differential equation depending only on the  $z$ -coordinate:

$$\Psi''(z) = \kappa_D^2 \sinh \Psi(z), \quad (6.18)$$

where  $\Psi \equiv \beta e \psi$  is the rescaled dimensionless potential and we have assumed that the external charge,  $\rho_f$ , is restricted to the system boundaries and will only affect the boundary conditions.

We will consider two boundary conditions in this section. A fixed surface potential (Dirichlet boundary condition),  $\Psi_s \equiv \Psi(z=0) = \text{const}$ , and constant surface charge (Neumann boundary condition),  $\sigma \propto \Psi'_s = \text{const}$ . A third and more specialized boundary condition of charge regulation will be treated in detail in Section 6.8. In the constant charge case, the membrane charge is modelled via a fixed surface charge density,  $\rho_f = \sigma \delta(z)$  in Eq. (6.8). A variation of the Helmholtz free energy,  $F$  of Eq. (6.11) with respect to the surface potential,  $\Psi_s$ ,  $\delta F / \delta \Psi_s = 0$ , is equivalent to constant surface charge boundary:

$$\left. \frac{d\Psi}{dz} \right|_{z=0} = -\frac{4\pi \ell_B \sigma}{e}. \quad (6.19)$$

Although we focus in the rest of the chapter on monovalent electrolytes, the extension to multivalent electrolytes is straightforward.

The boundary condition of Eq. (6.19) is valid if the electric field does not penetrate the ‘oily’ part of the membrane. This assumption can be justified (Kiometzis and Kleinert 1989, Winterhalter and Helfrich 1992), as long as  $\epsilon_L / \epsilon_w \approx 1/40 \ll h / \lambda_D$ , where  $h$  is the membrane thickness (see Figure 6.1). All our results for one or two flat membranes, Sections 6.3–6.4 and 6.5–6.7, respectively, rely on this decoupled limit where the two sides (monolayers) of the membrane are completely decoupled and the electric field inside the membrane is negligible.

#### 6.3.1 COUNTERIONS ONLY

A single charged membrane in contact with a cloud of counterions in solution is one of the simplest problems that has an analytical solution. It has been formulated and solved in

the beginning of the 20th century by Gouy (1910, 1917) and Chapman (1913). The aim is to find the profile of the counterion cloud forming a diffusive electric double layer close to a planar membrane (placed at  $z = 0$ ) with fixed surface charge density (per unit area),  $\sigma$ , as in Figure 6.2.

Without loss of generality, the single-membrane problem is treated here for negative (anionic) surface charges ( $\sigma < 0$ ) and positive monovalent counterions (cations) in the solution,  $q_+ = e$  and  $n(z) = n_+(z)$ , such that the charge neutrality condition

$$\sigma = -e \int_0^{\infty} n(z) dz \quad (6.20)$$

is fulfilled.

The PB equation for monovalent counterions is written as

$$\Psi''(z) = -4\pi\ell_B n_0 e^{-\Psi(z)}, \quad (6.21)$$

where  $n_0$  is the reference density, taken at zero potential in the absence of a salt reservoir. The PB equation, Eq. (6.21), with the boundary condition for one charged membrane, Eq. (6.19), and vanishing electric field at infinity, can be integrated analytically twice, yielding

$$\Psi(z) = 2 \ln(z + \ell_{GC}) + \Psi_0, \quad (6.22)$$

so that the density is

$$n(z) = \frac{1}{2\pi\ell_B} \frac{1}{(z + \ell_{GC})^2}, \quad (6.23)$$

where  $\Psi_0$  is a reference potential and  $\ell_{GC}$  is the Gouy–Chapman length defined in Eq. (6.2). For example, for a

choice of  $\Psi_0 = -2 \ln(\ell_{GC})$ , the potential at  $z = 0$  vanishes and Eq. (6.22) reads

$$\Psi(z) = 2 \ln(1 + z/\ell_{GC}). \quad (6.24)$$

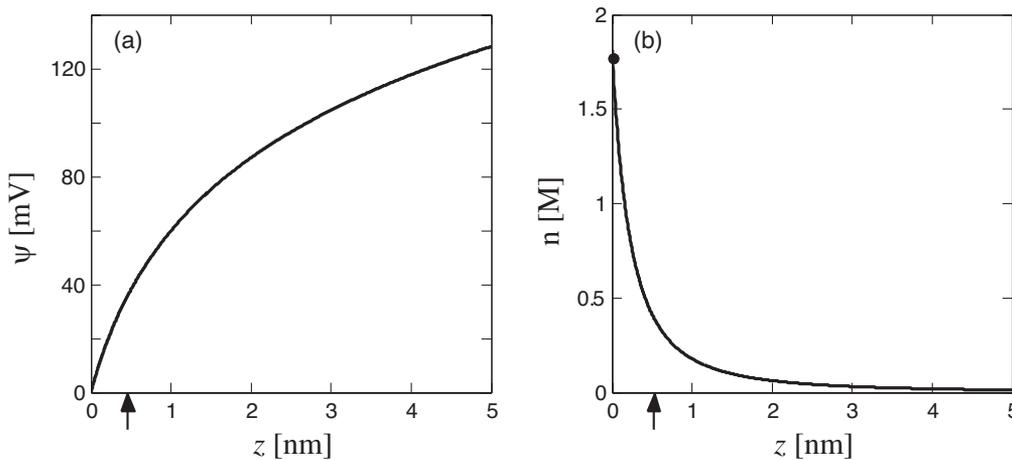
Although the entire counterion profile is diffusive as it decays algebraically, half of the counterions ( $|\sigma|/2$  per unit area) accumulate in a layer of thickness  $\ell_{GC}$  close to the membrane,

$$e \int_0^{\ell_{GC}} n(z) dz = \frac{1}{2} |\sigma|. \quad (6.25)$$

As an example, we present in Figure 6.3 the potential  $\psi$  (in mV) and ionic profile  $n$  (in M) for a surface density of  $\sigma = -e/(2 \text{ nm}^2)$ , leading to a Gouy–Chapman length,  $\ell_{GC} \approx 0.46 \text{ nm}$ . The figure clearly shows the build-up of the diffusive layer of counterions attracted by the negatively charged membrane, reaching a limiting value of  $n_s = n(0) \approx 1.82 \text{ M}$ . Note that the potential has a weak logarithmic divergence as  $z \rightarrow \infty$ . This divergence is a consequence of the vanishing ionic reservoir (counterions only) with counterion density obeying the Boltzmann distribution. However, the physically measured electric field,  $E = -d\psi/dz$ , properly decays to zero as  $\sim 1/z$ , at  $z \rightarrow \infty$ .

### 6.3.2 ADDED ELECTROLYTE

Another case of experimental interest is that of a single charged membrane at  $z = 0$  in contact with an electrolyte reservoir. For a symmetric electrolyte,  $n_+^{(b)} = n_-^{(b)} \equiv n_b$ , and the same boundary condition of constant surface charge  $\sigma$ , Eq. (6.19), holds at the  $z = 0$  surface. The negatively charged membrane attracts the counterions and repels the co-ions.



**FIGURE 6.3** The electric double layer for a single charged membrane in contact with an aqueous solution of neutralizing monovalent counterions. In (a), the electrostatic potential  $\psi(z)$  (in mV) is plotted as function of the distance from the membrane,  $z$ , Eq. (6.24). The charged membrane is placed at  $z=0$  with  $\sigma = -e/(2 \text{ nm}^2) < 0$ . The zero of the potential is chosen to be at the membrane,  $\psi(z=0)=0$ . In (b), the density profile of the counterions,  $n$  (in M), is plotted as function of the distance  $z$ . Its value at the membrane is  $n(z=0) = n_s \approx 1.82 \text{ M}$  and the Gouy–Chapman length,  $\ell_{GC} \approx 0.46 \text{ nm}$ , is marked by an arrow.

As will be shown below, the potential decays to zero from below at large  $z$ ; hence, it is always negative. Since the potential is a monotonic function, this also implies that  $\Psi'(z)$  is always positive. At large  $z$ , where the potential decays to zero, the ionic profiles tend to their bulk (reservoir) densities,  $n^\pm|_{\infty} = n_b$ .

The PB equation for monovalent electrolyte, Eq. (6.18), with the boundary conditions as explained above can be solved analytically. The first integration of the PB equation for 1:1 electrolyte yields

$$\frac{d\Psi}{dz} = -2\kappa_D \sinh(\Psi/2), \quad (6.26)$$

where we have used  $d\Psi/dz(z \rightarrow \infty) = 0$  that is implied by the Gauss law and electroneutrality, and chose the bulk potential,  $\Psi(z \rightarrow \infty) = 0$ , as the reference potential. A further integration yields

$$\Psi = -4 \tanh^{-1}(\gamma e^{-\kappa_D z}) = -2 \ln \left( \frac{1 + \gamma e^{-\kappa_D z}}{1 - \gamma e^{-\kappa_D z}} \right), \quad (6.27)$$

where  $\gamma$  is an integration constant,  $0 < \gamma < 1$ . Its value is determined by the boundary condition at  $z = 0$ .

The two ionic profiles,  $n_\pm(z)$ , are calculated from the Boltzmann distribution, Eq. (6.5), and from Eq. (6.27), yielding:

$$n_\pm(z) = n_b \left( \frac{1 \pm \gamma e^{-\kappa_D z}}{1 \mp \gamma e^{-\kappa_D z}} \right)^2. \quad (6.28)$$

For constant surface charge, the parameter  $\gamma$  is obtained by substituting the potential from Eq. (6.27) into the boundary

condition at  $z = 0$ , Eq. (6.19). This yields a quadratic equation,  $\gamma^2 + 2\kappa_D \ell_{GC} \gamma - 1 = 0$ , with  $\gamma$  as its positive root:

$$\gamma = -\kappa_D \ell_{GC} + \sqrt{(\kappa_D \ell_{GC})^2 + 1}. \quad (6.29)$$

For constant surface potential, the parameter  $\gamma$  can be obtained by setting  $z = 0$  in Eq. (6.27),

$$\Psi_s = e\psi_s/k_B T = -4 \tanh^{-1} \gamma. \quad (6.30)$$

We use the fact that the surface potential  $\Psi_s$  is uniquely determined by the two lengths,  $\ell_{GC} \sim \sigma^{-1}$  and  $\lambda_D$ , and write the electrostatic potential as

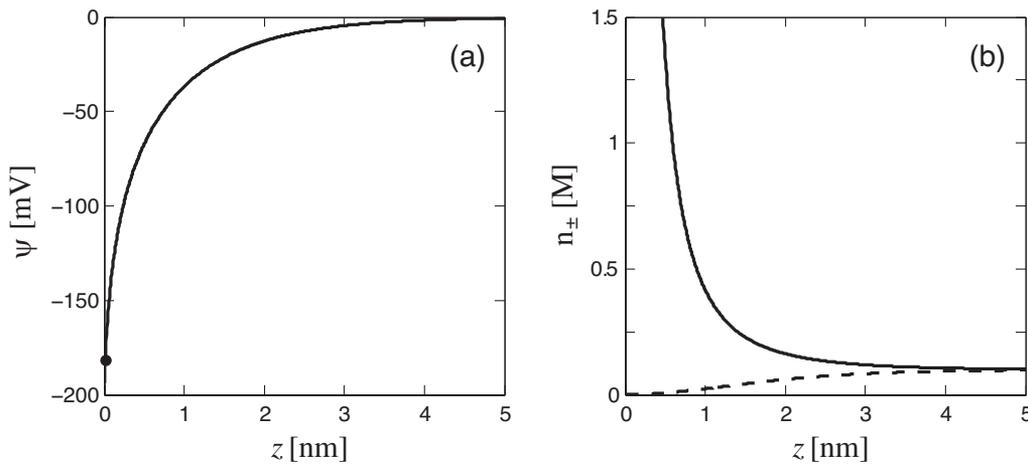
$$\Psi(z) = -2 \ln \left[ \frac{1 - \tanh(\Psi_s/4) e^{-\kappa_D z}}{1 + \tanh(\Psi_s/4) e^{-\kappa_D z}} \right], \quad (6.31)$$

where  $\Psi_s < 0$ , in accord with our choice of  $\sigma < 0$ . In Figure 6.4, we show typical profiles for the electrostatic potential and ionic densities, for  $\sigma = -5e/nm^2$  ( $\ell_{GC} \approx 0.046$  nm). Note that this surface charge density is ten times larger than  $\sigma$  of Figure 6.3. For electrolyte bulk density of  $n_b = 0.1$  M, the Debye screening length is  $\lambda_D \approx 0.97$  nm.

The DH (linearized) limit of the PB equation, Eq. (6.14), is obtained for small surface charge and/or high electrolyte strength,  $\kappa_D \ell_{GC} \gg 1$ . This limit yields  $\gamma \approx (2\kappa_D \ell_{GC})^{-1}$ , and the potential can be approximated as

$$\Psi \approx \Psi_s e^{-\kappa_D z} \approx -\frac{2}{\kappa_D \ell_{GC}} e^{-\kappa_D z}. \quad (6.32)$$

As expected for the DH limit, the solution is exponentially screened and falls off to zero for  $z \gg \kappa_D^{-1} = \lambda_D$ .



**FIGURE 6.4** The electric double layer for a single charged membrane in contact with a 1:1 monovalent electrolyte reservoir of concentration  $n_b = 0.1$  M, corresponding to  $\lambda_D \approx 0.97$  nm. The membrane located at  $z=0$  is negatively charged with  $\sigma = -5e/nm^2$ , yielding  $\ell_{GC} \approx 0.046$  nm. Note that the value of  $\sigma$  is ten times larger than the value used in Figure 6.3. In (a), we plot the electrostatic potential,  $\psi(z)$  as function of  $z$ , the distance from the membrane. The value of the surface potential is  $\psi_s \approx -194$  mV. In (b), the density profiles of counterions (solid line) and co-ions (dashed line),  $n_\pm$  (in M), are plotted as function of the distance from the membrane,  $z$ . The positive counterion density at the membrane is  $n_+(z=0) \approx 182$  M (not shown in the figure).

The counterion-only case, considered earlier in Section 6.3.1, is obtained by formally taking the  $n_b \rightarrow 0$  limit in Eqs. (6.27)–(6.29) or, equivalently,  $\kappa_D \ell_{GC} \ll 1$ . This means that  $\gamma \approx 1 - \kappa_D \ell_{GC}$ , and from Eq. (6.27), we recover Eq. (6.23) for the counterion density,  $n(z) = n_+(z)$ , while the co-ion density,  $n_-(z)$ , vanishes.

For a system in contact with an electrolyte reservoir, the potential always has an exponentially screened form in the distal region (far from the membrane). This can be seen by taking  $z \rightarrow \infty$  while keeping  $\kappa_D \ell_{GC}$  finite in Eq. (6.27):

$$\Psi(z) \approx -4\gamma e^{-\kappa_D z}. \quad (6.33)$$

Moreover, it is possible to extract from the distal form an effective surface charge density,  $\sigma_{\text{eff}}$ , by comparing the coefficient  $4\gamma$  of Eq. (6.33) with an effective coefficient  $2/(\kappa_D \ell_{GC})$  from the DH form, Eq. (6.32),

$$|\sigma_{\text{eff}}| = 2\gamma \kappa_D \ell_{GC} |\sigma| = \frac{e \kappa_D}{\pi \ell_B} \gamma. \quad (6.34)$$

Note that  $\gamma = \gamma(\kappa_D \ell_{GC})$  is calculated for the nominal parameter values in Eq. (6.29). The same concept of an effective  $\sigma$  is useful in several situations other than the simple planar geometry considered here.

### 6.3.3 THE GRAHAME EQUATION

In the planar geometry, for any amount of salt, the nonlinear PB equation can be integrated analytically, resulting in a useful relation known as the Grahame equation (Grahame 1947). This equation is a relation between the surface charge density,  $\sigma$ , and the limiting value of the ionic density profile at the membrane,  $n_{\pm}^{(s)} \equiv n_{\pm}(z=0)$ . The first integration of the PB equation for a 1:1 electrolyte yields, Eq. (6.26),  $d\Psi/dz = -2\kappa_D \sinh(\Psi/2)$ . Using the boundary condition, Eq. (6.19), and simple hyperbolic function identities gives a relation between  $\sigma$  and  $\Psi_s$

$$\pi \ell_B \left( \frac{\sigma}{e} \right)^2 = n_b \left( \cosh(\Psi_s) - 1 \right), \quad (6.35)$$

and via the Boltzmann distribution of  $n_{\pm}$ , the Grahame equation is obtained:

$$\sigma^2 = \frac{e^2}{2\pi \ell_B} \left( n_+^{(s)} + n_-^{(s)} - 2n_b \right). \quad (6.36)$$

This equation implies a balance of stresses on the surface, with the Maxwell stress of the electric field compensating the van 't Hoff ideal pressure of the ions.

For large and negative surface potential,  $|\Psi_s| \gg 1$ , the co-ion density,  $n_-^{(s)} \sim \exp(-|\Psi_s|)$ , can be neglected, and Eq. (6.36) becomes

$$\sigma^2 = \frac{e^2}{2\pi \ell_B} \left( n_+^{(s)} - 2n_b \right). \quad (6.37)$$

For example, for a surface charge density of  $\sigma = -5e/nm^2$  (as in Figure 6.4) and an ionic strength of  $n_b = 0.1M$ , the limiting value of the counterion density at the membrane is  $n_+^{(s)} \approx 182M$  and that of the co-ions is  $n_-^{(s)} \approx 5 \times 10^{-5}M$ . The very high and unphysical value of  $n_+^{(s)}$  should be understood as an artefact of the continuum PB theory. In physical situations, the ions accumulate in the membrane vicinity till their concentration saturates due to the finite ionic size and other ion–surface interactions. We will further explore this point in Sections 6.4 and 6.8.

The differential capacitance is another useful quantity to calculate, and it gives a physical measurable surface property. By using Eq. (6.35), we obtain

$$C_{PB} = \frac{d\sigma}{d\Psi_s} = \frac{e}{k_B T} \frac{d\sigma}{d\Psi_s} = \epsilon_0 \epsilon_w \kappa_D \cosh(\Psi_s/2). \quad (6.38)$$

As shown later in Figure 6.6, the PB differential capacitance has a minimum at the potential of zero charge,  $\Psi_s = 0$ , and increases exponentially for  $|\Psi_s| \gg 1$ .

## 6.4 MODIFIED POISSON–BOLTZMANN THEORY

Within the PB theory, the density of accumulated counterions at the membrane might reach unphysical high values (see Figure 6.4). This unphysical situation is avoided on the MF level by accounting for the solvent entropy. Including this additional term yields a modified free-energy and PB equation (mPB). Taking this entropy into account yields a modified free energy, written here for monovalent electrolyte:

$$\begin{aligned} \beta F = \int_V d^3r \left[ -\frac{1}{8\pi \ell_B} |\nabla \Psi(\mathbf{r})|^2 + [n_+(\mathbf{r}) - n_-(\mathbf{r})] \Psi(\mathbf{r}) \right. \\ \left. + n_+ \ln(n_+ + a^3) + n_- \ln(n_- - a^3) \right. \\ \left. + \frac{1}{a^3} (1 - a^3 n_+ - a^3 n_-) \ln(1 - a^3 n_+ - a^3 n_-) \right]. \end{aligned} \quad (6.39)$$

This is the free energy of a Coulomb lattice gas (Borukhov, Andelman and Orland 1997, 2000, Kilic, Bazant and Ajdari 2007, Adar et al 2018). Taking the variation of the above free energy with respect to  $n_{\pm}$ ,  $\delta F / \delta n_{\pm} = \mu_{\pm}$ , gives the ionic profiles

$$n_{\pm}(z) = \frac{n_b e^{\mp \Psi}}{1 - 2\phi_b + 2\phi_b \cosh \Psi}, \quad (6.40)$$

with  $\phi_b = n_b a^3$  being the bulk volume fraction of the ions. For simplicity,  $a$  is taken to be the same molecular size of all ionic species and the solvent.

### ENTROPY DERIVATION OF THE MPB

Let us start with a homogenous system containing an ionic solution inside a volume  $v$ , with  $N_+$  cations,  $N_-$  anions and  $N_w$  water molecules, such that  $N_+ + N_- + N_w = N$ . The number of different combinations of cations, anions, and water molecules is  $N!/(N_+!N_-!N_w!)$ . Therefore, the entropy is

$$S_v = -k_B \log \left( \frac{N!}{N_+!N_-!N_w!} \right) \approx -k_B \left[ N_+ \ln \left( \frac{N_+}{N} \right) + N_- \ln \left( \frac{N_-}{N} \right) + (N - N_+ - N_-) \ln \left( 1 - \frac{N_+}{N} - \frac{N_-}{N} \right) \right], \quad (6.41)$$

where we have used Stirling's formula for  $N_{\pm}, N_w \gg 1$ .

We now consider a system of volume  $V \gg v$ . The entropy of such system can be written in the continuum limit as

$$S_V = \int \frac{d^3r}{v} S_v = k_B \int d^3r \left[ n_+ \ln(n_+ a^3) + n_- \ln(n_- a^3) + \frac{1}{a^3} (1 - a^3 n_+ - a^3 n_-) \ln(1 - a^3 n_+ - a^3 n_-) \right]. \quad (6.42)$$

where  $n_{\pm} = N_{\pm}/v$  and  $n_w = N_w/v$  are the densities of the cations, anions, and water molecules, respectively, and  $N = v/a^3$  is the total number of molecules in the volume  $v$ . In this last equation, we have used the lattice gas formulation, in which the solution is modelled as a cubic lattice with unit cell of size  $a \times a \times a$ . Each unit cell contains only one molecule,  $a^3(n_+ + n_- + n_w) = 1$ .

In the above equation, we have also used the equilibrium relation

$$e^{\beta\mu_{\pm}} = \frac{n_{\pm} a^3}{1 - 2n_{\pm} a^3} = \frac{\phi_b}{1 - 2\phi_b}, \quad (6.43)$$

valid in the bulk where  $\Psi = 0$ . Variation with respect to  $\Psi$ ,  $\delta F/\delta \Psi = 0$ , yields the mPB equation for 1:1 electrolyte:

$$\nabla^2 \Psi(\mathbf{r}) = -4\pi\ell_B [n_+(\mathbf{r}) - n_-(\mathbf{r})] = \frac{\kappa_D^2 \sinh \Psi}{1 - 2\phi_b + 2\phi_b \cosh \Psi}. \quad (6.44)$$

For small electrostatic potentials,  $|\Psi| \ll 1$ , the ionic distribution, Eq. (6.40), reduces to the usual Boltzmann distribution, but for large electrostatic potentials,  $|\Psi| \gg 1$ , this model gives very different results with respect to the PB theory. In particular, the ionic concentration is unbound in the standard PB theory, whereas it is bound for the mPB by the close packing density,  $1/a^3$ . This effect is important close to strongly charged membranes immersed in an electrolyte solution, while the regular PB equation is recovered in the dilute bulk limit,  $n_{\pm} a^3 \ll 1$ , for which the solvent entropy can be neglected.

For large electrostatic potentials, the contribution of the co-ions is negligible, and the counterion concentration follows a distribution reminiscent of the Fermi–Dirac distribution

$$n_-(\mathbf{r}) \approx \frac{1}{a^3} \frac{1}{1 + \exp(-\Psi - \beta\mu)}, \quad (6.45)$$

where electroneutrality dictates  $\mu = \mu_{\pm}$ . In Figure 6.5, we show for comparison the modified and regular PB profiles

for a 1:1 electrolyte. To emphasize the saturation effect of the mPB theory, we chose in the figure a large ion size,  $a = 0.8 \text{ nm}$ .

The mPB theory also implies a modified Grahame equation that relates the surface charge density to the ion surface density,  $n_{\pm}^{(s)}$ . First, we find the relation between  $\sigma$  and the surface potential,  $\Psi_s$ ,

$$\left( \frac{\sigma}{e} \right)^2 = \frac{1}{2\pi a^3 \ell_B} \ln \left[ 1 + 2\phi_b (\cosh(\Psi_s) - 1) \right]. \quad (6.46)$$

This equation represents a balance of stresses on the surface, where the Maxwell stress of the electric field is equal to the lattice-gas pressure of the ions. The surface potential can also be calculated:

$$\Psi_s = \cosh^{-1} \left( \frac{e^{\xi} - 1 + 2\phi_b}{2\phi_b} \right), \quad (6.47)$$

with the dimensionless parameter  $\xi = a^3 / (2\pi\ell_B \ell_{GC}^2)$ .

For large surface charge or large surface potential, the co-ions concentration at the membrane is negligible,  $n_{\pm}^{(s)} \ll 1$ , and the surface potential, Eq. (6.47) is approximated by

$$\Psi_s \approx \ln(e^{\xi} - 1 + 2\phi_b) - \ln(\phi_b), \quad (6.48)$$

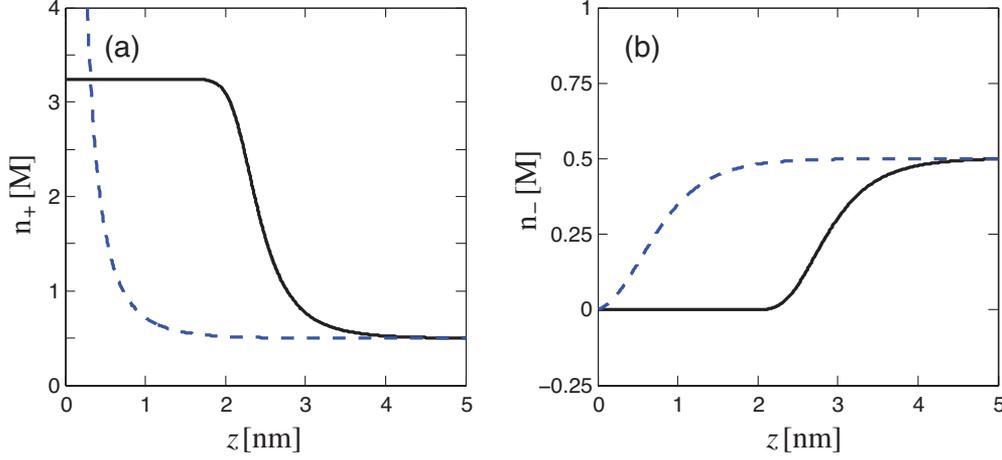
and from Eq. (6.46), we obtain the Grahame equation,

$$\left( \frac{\sigma}{e} \right)^2 \approx \frac{1}{2\pi a^3 \ell_B} \ln \left( \frac{1 - 2\phi_b}{1 - a^3 n_{\pm}^{(s)}} \right). \quad (6.49)$$

Note that in the dilute limit  $\phi_b \ll 1$ , the Grahame equation reduces to the regular PB case (Eq. 6.36).

1. See last page:  
sign error

2. See last page:  
sign error



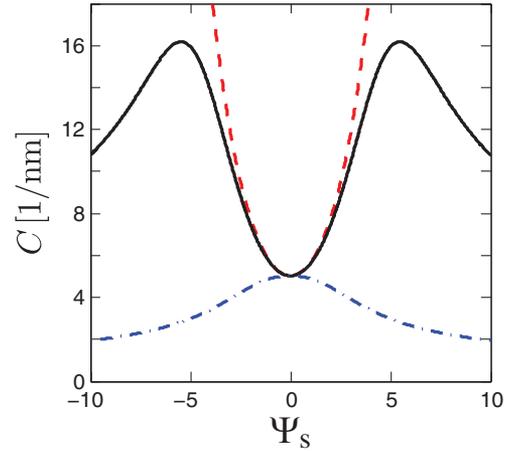
**FIGURE 6.5** Comparison of the modified PB (mPB) profiles (black solid lines) with those of the regular PB one (dashed blue lines). In (a), we show the counterion profile, and in (b), the co-ion profile. The parameters used are ion size  $a=0.8$  nm, surface charge density  $\sigma=-5e/\text{nm}^2$ , and 1:1 electrolyte with ionic strength  $n_b=0.5$  M. Note that while the PB value at the membrane is  $n_s^+ \approx 182$  M, the mPB density saturates at  $n_s^+ \approx 3.2$  M.

It is also straightforward but more cumbersome to calculate the differential capacitance,  $C = d\sigma/d\Psi_s$ , for the mPB theory. From Eq. (6.46), we obtain (Kornyshev 2007, Nakayama and Andelman 2015)

$$C_{\text{mPB}} = \frac{C_{\text{PB}}}{1 + 4\phi_b \sinh^2(\Psi_s/2)} \sqrt{\frac{4\phi_b \sinh^2(\Psi_s/2)}{\ln[1 + 4\phi_b \sinh^2(\Psi_s/2)]}}. \quad (6.50)$$

Although it can be shown that for  $\phi_b \rightarrow 0$  the mPB differential capacitance reduces to the standard PB result, the resulting  $C_{\text{mPB}}$  is quite different for any finite value of  $\phi_b$ . The main difference is that instead of an exponential divergence of  $C_{\text{PB}}$  at large potentials,  $C_{\text{mPB}}$  decreases for high-bias  $|\Psi_s| \gg 1$ . For rather small bulk densities,  $\phi_b < 1/6$ , the  $C_{\text{mPB}}$  shows a behaviour called *camel-shape* or *double-hump*. This behaviour is also observed in experiments at relatively low salt concentrations. As shown in Figure 6.6, the double-hump  $C_{\text{mPB}}$  has a minimum at  $\Psi_s = 0$  and two maxima. The peak positions can roughly be estimated by substituting the close packing concentration,  $n = 1/a^3$ , into the Boltzmann distribution (Eq. 6.5), yielding  $\Psi_s^{\text{max}} \approx \mp \ln(\phi_b)$ . Using parameter values as in Figure 6.6,  $\Psi_s^{\text{max}}$  is estimated as  $\pm 4.6$  as compared with the exact values,  $\Psi_s^{\text{max}} = \pm 5.5$ .

Furthermore, it can be shown that for high salt densities,  $\phi_b > 1/6$ ,  $C_{\text{mPB}}$  exhibits (see also Figure 6.6) a unimodal maximum close to the potential of zero charge, rather than a minimum as for  $C_{\text{PB}}$ . Such results that take into account finite ion size for the differential capacitance are of importance in the theory of confined ionic liquids (Kornyshev 2007, Nakayama and Andelman 2015).



**FIGURE 6.6** Comparison of the differential capacitance,  $C$ , calculated from the regular PB theory (dashed red line), Eq. (6.38), with  $n_b \approx 0.4$  mM (chosen so that it corresponds to  $\phi_b=0.01$  and  $a=0.3$  nm), and from the mPB theory, Eq. (6.50). The mPB differential capacitance is calculated for  $a=0.3$  nm. For low  $\phi_b=0.01$ , it shows a camel shape (black solid line), while for high  $\phi_b=0.2$ , it shows a unimodal (dash-dotted blue line).

## 6.5 TWO-MEMBRANE SYSTEM: OSMOTIC PRESSURE

We now consider the PB theory of two charged membranes as shown in Figure 6.7. The two membranes can, in general, have different surface charge densities:  $\sigma_1$  at  $z = -d/2$  and  $\sigma_2$  at  $z = d/2$ . The boundary conditions of the two-membrane system are written as  $\rho_f = \sigma_1\delta(z+d/2) + \sigma_2\delta(z-d/2)$ , and using the variation of the free energy,  $\delta F/\delta\Psi_s = 0$ :

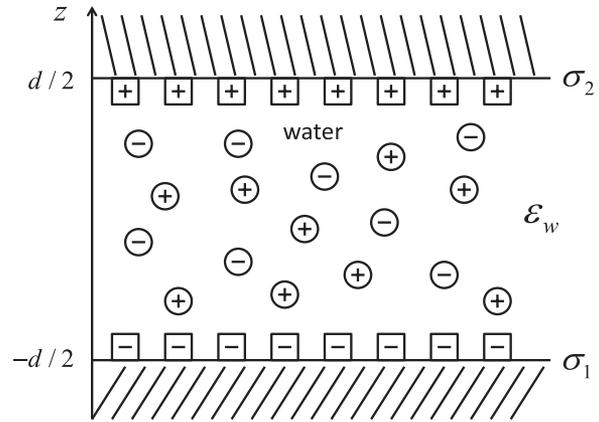
$$\begin{aligned} \Psi'|_{-d/2} &= -4\pi\ell_B \frac{\sigma_1}{e}, \\ \Psi'|_{d/2} &= 4\pi\ell_B \frac{\sigma_2}{e}. \end{aligned} \quad (6.51)$$

It is of interest to calculate the force (or the osmotic pressure) between two membranes interacting across the ionic solution. The osmotic pressure is defined as  $\Pi = P_{\text{in}} - P_{\text{out}}$ , where  $P_{\text{in}}$  is the inner pressure and  $P_{\text{out}}$  is the pressure exerted by the reservoir that is in contact with the two-membrane system. Sometimes the osmotic pressure is referred to as the disjoining pressure, introduced first by Derjaguin (Churaev, Derjaguin and Muller 2014).

Let us start by calculating the inner and outer pressures from the Helmholtz free energy. The pressure ( $P_{\text{in}}$  or  $P_{\text{out}}$ ) is the variation of the free energy with the volume:

$$P = -\frac{\delta F}{\delta V} = -\frac{1}{A} \frac{\delta F}{\delta d}, \quad (6.52)$$

with  $V = Ad$  being the system volume,  $A$  the lateral membrane area, and  $d$  the intermembrane distance. As the interaction between the two membranes can be either attractive ( $\Pi < 0$ ) or repulsive ( $\Pi > 0$ ), we will analyse the criterion for the crossover ( $\Pi = 0$ ) between these two regimes as function of the surface charge asymmetry and intermembrane distance.



**FIGURE 6.7** Schematic drawing of two asymmetric membranes. The planar membrane located at  $z = -d/2$  carries a charge density  $\sigma_1$ , while the membrane at  $z = d/2$  has a charge density of  $\sigma_2$ . The antisymmetric membrane set-up is a special case with  $\sigma_1 = -\sigma_2$ , while in the symmetric case,  $\sigma_1 = \sigma_2$ .

### GENERAL DERIVATION OF THE PRESSURE

The Helmholtz free-energy obtained from Eq. (6.11) can be written in a general form as  $F = A \int f[\Psi(z), \Psi'(z)] dz$ , where we use the Poisson equation to obtain the relation,  $n_{\pm} = n_{\pm}(\Psi)$ . As the integrand  $f$  depends only implicitly on the  $z$  coordinate through  $\Psi(z)$ , one can obtain from the Euler–Lagrange equations the following relation (Ben-Yaakov et al. 2009a):

$$\begin{aligned} f - \sum_i \mu_i^{\text{tot}} n_i - \frac{\partial f}{\partial \Psi'} \Psi' &= \text{const} \\ &= \frac{k_B T}{8\pi\ell_B} (\Psi')^2 + k_B T \sum_{i=1}^M z_i n_i \Psi + k_B T \sum_{i=1}^M [n_i \ln(n_i a^3) - n_i], \end{aligned} \quad (6.53)$$

where the sum is over  $i = 1, \dots, M$  ionic species and the total chemical potential defined as before,  $\delta F / \delta n_i = \partial f / \partial n_i = \mu_i^{\text{tot}}$ . Let us understand the meaning of the constant on the right-hand side of the above equation. For uncharged solutions, the Helmholtz free energy per unit volume contains only the entropy term,  $f = k_B T \sum_i [n_i \ln(n_i a^3) - n_i]$ , and from Eq. (6.53), we obtain  $f - \sum_i \mu_i^{\text{tot}} n_i = \text{const}$ . A known thermodynamic relation is  $P = \sum_i \mu_i^{\text{tot}} n_i - f$ , implying that the right-hand side constant is  $-P$ . However, even for charged liquid mixtures, the electrostatic potential vanishes in the bulk, away from the boundaries, and reduces to the same value as for uncharged solutions. Therefore, we conclude that the right-hand side constant is  $-P$ , yielding

$$P = -\frac{k_B T}{8\pi\ell_B} (\Psi')^2 + k_B T \sum_{i=1}^M n_i. \quad (6.54)$$

If the electric field and ionic densities are calculated right at the surface, we obtain the *contact theorem* that gives the osmotic pressure acting on the surface. Another and more straightforward way to calculate the pressure is to calculate the incremental difference in free energy,  $F$ , for an intermembrane separation  $d$ , i.e.,  $[F(d + \delta d) - F(d)] / \delta d$ . The calculation of  $F(d + \delta d)$  can be done by including an additional slab of width  $\delta d$  in the space between the two membranes at an arbitrary position. We remark that the validity of the contact theorem itself is not limited to the PB theory but is an exact theorem of statistical mechanics (Henderson and Blum 1981, Evans and Wennerström 1999, Dean and Horgan 2003).

We are interested in the osmotic pressure,  $\Pi$ . For an ionic reservoir in the dilute limit, Eq. (6.54) gives  $P_{\text{out}} = k_B T \sum_i n_i^{(b)}$ , where  $n_i^{(b)}$  is the  $i^{\text{th}}$  ionic species bulk density. Thus, the osmotic pressure can be written as

$$\Pi = -\frac{k_B T}{8\pi\ell_B} (\Psi')^2 + k_B T \sum_{i=1}^M (n_i(z) - n_i^{(b)}) \quad (6.55)$$

and for monovalent 1:1 ions:

$$\Pi = -\frac{k_B T}{8\pi\ell_B} (\Psi')^2 + 2k_B T n_b (\cosh[\Psi(z)] - 1) = \text{const.} \quad (6.56)$$

At any position  $z$  between the membranes, the osmotic pressure has two contributions. The first is a negative Maxwell electrostatic pressure proportional to  $(\Psi')^2$ . The second is due to the entropy of mobile ions and measures the local entropy change (at an arbitrary position,  $z$ ) with respect to the ion entropy in the reservoir.

## 6.6 TWO SYMMETRIC MEMBRANES, $\sigma_1 = \sigma_2$

For two symmetrically charged membranes,  $\sigma_1 = \sigma_2 \equiv \sigma$  at  $z = \pm d/2$ , the electrostatic potential is symmetric about the midplane yielding a zero electric field,  $E = 0$  at  $z = 0$ . It is then sufficient to consider the interval  $[0, d/2]$  with the boundary conditions,

$$\Psi'|_{z=d/2} = \Psi'_s = 4\pi\ell_B \sigma / e, \quad \Psi'|_{z=0} = \Psi'_m = 0. \quad (6.57)$$

As  $\Pi$  is constant (independent of  $z$ ) between the membranes, one can calculate the disjoining pressure,  $\Pi$ , from Eq. (6.55), at any position  $z$ , between the membranes. A simple choice will be to evaluate it at  $z = 0$  (the midplane), where the electric field vanishes for the symmetric  $\sigma_1 = \sigma_2$  case,

$$\Pi = k_B T \sum_{i=1}^M (n_i^{(m)} - n_i^{(b)}) = k_B T \sum_{i=1}^M n_i^{(b)} (e^{-z\Psi_m} - 1) > 0, \quad (6.58)$$

and for monovalent ions,  $z_i = \pm 1$ , we get

$$\Pi = 4k_B T n_b \sinh^2(\Psi_m/2) > 0, \quad (6.59)$$

where  $n_i^{(m)} = n_i(z=0)$  is the midplane concentration of the  $i^{\text{th}}$  species. It can be shown that the electroneutrality condition implies that the osmotic pressure is always *repulsive* for any shape of boundaries (Sader and Chan 1999, Neu 1999) as long as we have two symmetric membranes ( $\sigma_1 = \sigma_2$ ).

Note that the Grahame equation can be derived also for the two-membrane case with added electrolyte. One way of doing it is by comparing the pressure of Eq. (6.55)

evaluated at one of the membranes,  $z = \pm d/2$ , and at the midplane,  $z = 0$ . The pressure is constant between the two membranes; thus, by equating these two pressure expressions, the Grahame equation emerges

$$\left(\frac{\sigma}{e}\right)^2 = \frac{1}{2\pi\ell_B} \sum_{i=1}^M (n_i^{(s)} - n_i^{(m)}). \quad (6.60)$$

By taking the limit of infinite separation between the two membranes and  $n_i^{(m)} \rightarrow n_i^{(b)}$ , the Grahame equation for a single membrane (Eq. 6.36) is recovered.

### 6.6.1 COUNTERIONS ONLY

In the absence of an external salt reservoir, the only ions in the solution for a symmetric two-membrane system are positive monovalent ( $z = +1$ ) counterions with density  $n(z)$  that neutralizes the negative surface charge,

$$2\sigma = -e \int_{-d/2}^{d/2} n(z) dz. \quad (6.61)$$

The PB equation has an analytical solution for this case. Integrating twice the PB equation (Eq. 6.18) with the appropriate boundary conditions (Eq. 6.57) yields an analytical expression for the electrostatic potential:

$$\Psi(z) = \ln(\cos^2 Kz), \quad (6.62)$$

and consequently the counterion density is

$$n(z) = n_m e^{-\Psi(z)} = \frac{n_m}{\cos^2(Kz)}. \quad (6.63)$$

In the above, we have defined  $n_m = n(z=0)$  and chose arbitrarily  $\Psi_m = 0$ . We also introduced a new length scale,  $K^{-1}$ , related to  $n_m$  by

$$K^2 = 2\pi\ell_B n_m. \quad (6.64)$$

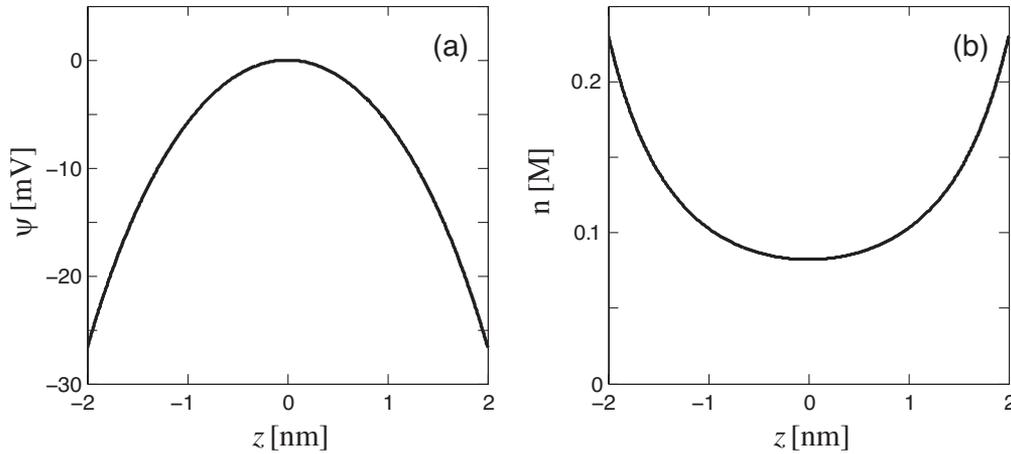
Notice that  $K$  plays a role similar to the inverse Debye length  $\kappa_D = \sqrt{8\pi\ell_B n_b}$ , with the midplane density replacing the bulk density,  $n_b \rightarrow n_m$ . Using the boundary condition at  $z = d/2$ , we get a transcendental relation for  $K$

$$Kd \tan(Kd/2) = \frac{d}{\ell_{GC}}. \quad (6.65)$$

In Figure 6.8, we show a typical counterion profile with its corresponding electrostatic potential for  $\sigma = -e/(7 \text{ nm}^2)$  and  $d = 4 \text{ nm}$ .

The osmotic pressure (Eq. 6.55), calculated for the counterion-only case, is

$$\Pi = \frac{k_B T}{2\pi\ell_B} K^2. \quad (6.66)$$



**FIGURE 6.8** The counterion-only case for two identically charged membranes located at  $z = \pm d/2$  with  $d = 4$  nm and  $\sigma = -e/(7 \text{ nm}^2)$  on each membrane ( $\ell_{\text{GC}} \simeq 1.6$  nm). In (a), we plot the electrostatic potential,  $\psi$ , and in (b), the counterion density profile,  $n$ . The plots are obtained from Eqs. (6.62)–(6.65).

For weak surface charge,  $d/\ell_{\text{GC}} \ll 1$ , one can approximate  $(Kd)^2 \simeq 2d/\ell_{\text{GC}} \ll 1$ , and the pressure is given by

$$\Pi \simeq -\frac{2k_{\text{B}}T\sigma}{e} \frac{1}{d} = \frac{k_{\text{B}}T}{\pi\ell_{\text{B}}\ell_{\text{GC}}} \frac{1}{d} \sim \frac{1}{d}. \quad (6.67)$$

The  $\Pi \sim 1/d$  behaviour is similar to an ideal-gas equation of state,  $P = Nk_{\text{B}}T/V$  with  $V = Ad$  and  $N$  the total number of counterions. The density (per unit volume) of the counterions is almost constant between the two membranes and is equal to  $2|\sigma|/(ed)$ . This density neutralizes the surface charge density,  $\sigma$ , on the two membranes. The main contribution to the pressure comes from an ideal-gas like pressure of the counterion cloud. This regime can be reached experimentally for small intermembrane separation,  $d < \ell_{\text{GC}}$ . For example, for  $e/\sigma$  in the range of 1–100 nm<sup>2</sup>,  $\ell_{\text{GC}} \sim 1/\sigma$  varies between 0.2 nm and 20 nm.

For the opposite case of strong surface charge,  $d/\ell_{\text{GC}} \gg 1$ , one gets  $Kd \simeq \pi$  from Eq. (6.65). This is the Gouy–Chapman regime. It is very different from the weak surface charge, as the density profile between the two membranes varies substantially leading to  $n_s \gg n_m$ , and to a pressure

$$\Pi \simeq \frac{\pi k_{\text{B}}T}{2\ell_{\text{B}}d^2} \sim \frac{1}{d^2}. \quad (6.68)$$

It is interesting to note that the above pressure expression does not depend explicitly on the surface charge density. This can be rationalized as follows. Counterions are accumulated close to the surface, at an average separation  $\ell_{\text{GC}} \sim 1/|\sigma|$ . Therefore, creating a surface dipole density of  $|\sigma|\ell_{\text{GC}}$ . The interaction energy per unit area is proportional to the electrostatic energy between two such planar dipolar layers, which scales as  $1/d$  for the free energy density and  $d^{-2}$  for the pressure. The surface charge density dependence itself vanishes because the effective dipolar-moment

surface density,  $|\sigma|\ell_{\text{GC}}$ , is charge independent. In the Gouy–Chapman regime, the electrostatic interactions are most dominant as they are long-ranged and unscreened. Of course, even in pure water, the effective Debye screening length is about 1  $\mu\text{m}$ , and the electrostatic interactions will be screened for larger distances.

### 6.6.2 ADDED ELECTROLYTE

When two charged membranes are placed in contact with an electrolyte reservoir, the co-ions and counterions between the membranes have a nonhomogenous density profile. The PB equation does not have a closed-form analytical solution for two (or more) ionic species, even when we restrict ourselves to a 1:1 symmetric and monovalent electrolyte. Instead, the solution can be expressed in terms of elliptic functions.

The PB equation for a monovalent 1:1 electrolyte (Eq. 6.18) is  $\Psi''(z) = \kappa_{\text{D}}^2 \sinh \Psi$ , while the same boundary conditions as in Eq. (6.57) are satisfied. The first integration from the midplane ( $z = 0$ ) to an arbitrary point between the membranes,  $z \in [-d/2, d/2]$ , gives

$$\frac{d\Psi}{dz} = -\kappa_{\text{D}} \sqrt{2 \cosh \Psi(z) - 2 \cosh \Psi_m}. \quad (6.69)$$

As explained in the beginning of Section 6.6,  $\Psi'_m = 0$  for two symmetric membranes and the second integration leads to an elliptic integral (see box below)

$$z = -\lambda_{\text{D}} \int_{\Psi_m}^{\Psi} \frac{d\eta}{\sqrt{2 \cosh \eta - 2 \cosh \Psi_m}}. \quad (6.70)$$

Inverting the relation  $z = z(\Psi)$  leads to the expression for the profile,  $\Psi(z)$ .

### THE ELECTROSTATIC POTENTIAL VIA JACOBI ELLIPTIC FUNCTIONS

It is possible to write Eq. (6.70) in terms of an incomplete elliptic integral of the first kind

$$F(\theta | a^2) \equiv \int_0^\theta \frac{d\eta}{\sqrt{1 - a^2 \sin^2 \eta}}. \quad (6.71)$$

After change of variables and some algebra, we write Eq. (6.70) with the help of Eq. (6.71) as

$$z = 2\lambda_D \sqrt{m} \left[ F\left(\frac{\pi}{2} \middle| m^2\right) - F(\varphi | m^2) \right], \quad (6.72)$$

with  $m = \exp(\Psi_m)$  and  $\varphi = \sin^{-1} \left[ \exp([\Psi - \Psi_m] / 2) \right]$ .

The electrostatic potential, which is the inverse relation of Eq. (6.72), can then be written in terms of the Jacobi elliptic function,  $\text{cd}(u | a^2)$ ,

$$\Psi = \Psi_m + 2 \ln \left[ \text{cd} \left( \frac{z}{2\lambda_D \sqrt{m}} \middle| m^2 \right) \right]. \quad (6.73)$$

In writing this equation, we have used the definition of the Jacobi elliptic functions:

$$\text{sn}(u | a^2) = \sin \alpha, \quad (6.74)$$

$$\text{cn}(u | a^2) = \cos \alpha = \sqrt{1 - \text{sn}^2(u | a^2)}, \quad (6.75)$$

$$\text{dn}(u | a^2) = \sqrt{1 - a^2 \text{sn}^2(u | a^2)}, \quad (6.76)$$

$$\text{cd}(u | a^2) = \frac{\text{cn}(u | a^2)}{\text{dn}(u | a^2)}, \quad (6.77)$$

with  $u \equiv F(\alpha | a^2)$ .

Using one of the boundary conditions (Eq. 6.57) with the first integration (Eq. 6.69) yields

$$\cosh(\Psi_s) = \cosh(\Psi_m) + 2 \left( \frac{\lambda_D}{\ell_{GC}} \right)^2. \quad (6.78)$$

The above equation also gives a relation between  $\sigma$  and the midplane potential,  $\Psi_m$ , in terms of Jacobi elliptic functions (see box above),

$$\frac{\sigma}{e} = \frac{\kappa_D}{4\pi\ell_B} \frac{m^2 - 1}{\sqrt{m}} \frac{\text{sn}(u_s | m^2)}{\text{cn}(u_s | m^2) \text{dn}(u_s | m^2)}, \quad (6.79)$$

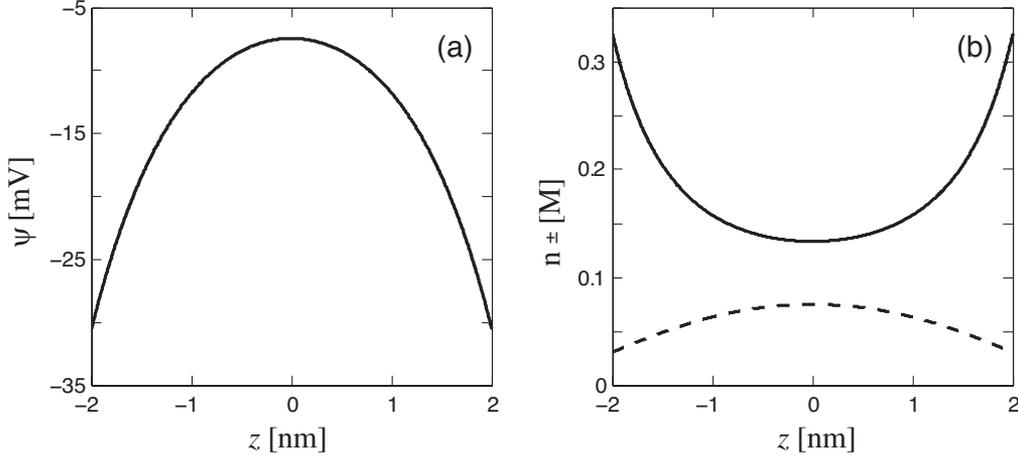
with  $u_s \equiv d / (4\lambda_D \sqrt{m})$  and  $m = \exp(\Psi_m)$  as defined after Eq. (6.72). For fixed surface charge, this relation gives the midplane potential,  $\Psi_m$ , and the osmotic pressure can then be calculated from Eq. (6.58). The other boundary condition can also be expressed as an elliptic integral

$$\begin{aligned} \frac{d}{2\lambda_D} &= - \int_{\Psi_m}^{\Psi_s} \frac{d\eta}{\sqrt{2 \cosh \eta - 2 \cosh \Psi_m}} \\ &= 2\sqrt{m} \left[ F\left(\frac{\pi}{2} \middle| m^2\right) - F(\varphi_s | m^2) \right], \end{aligned} \quad (6.80)$$

where  $\varphi_s = \sin^{-1} \left[ \exp([\Psi_s - \Psi_m] / 2) \right]$ .

The three equations, Eqs. (6.70), (6.78), and (6.80), completely determine the potential  $\Psi(z)$ , the two species density profiles,  $n_\pm(z) = n_b \exp(\mp\Psi)$ , and their midplane values  $n_\pm^{(m)} = n_b \exp(\mp\Psi_m)$ , as function of the three parameters: the intermembrane spacing  $d$ , the surface charge density  $\sigma$  (or equivalently  $\ell_{GC}$ ), and the electrolyte bulk ionic strength  $n_b$  (or equivalently  $\lambda_D$ ).

The exact form of the profiles and pressure can be obtained either from the numerical solution of Eqs. (6.70), (6.78) and (6.80) or by the usage of the elliptic functions. For example, we calculate numerically the counterion,



**FIGURE 6.9** Monovalent 1:1 electrolyte with  $n_b = 0.1$  M ( $\lambda_D \simeq 0.97$  nm) between two identically charged membranes with  $\sigma = -e/(7 \text{ nm}^2)$  each ( $\ell_{GC} \simeq 1.6$  nm) located at  $z = \pm d/2$  with  $d = 4$  nm. In (a), we plot the electrostatic potential,  $\psi$ , and in (b), we show the co-ion (dashed line) and counterion (solid line) density profiles,  $n_{\pm}$ . The plots are obtained from Eqs. (6.70), (6.78), and (6.80).

co-ion, and potential profiles as shown in Figure 6.9, where the three relevant lengths are  $d = 4$  nm,  $\ell_{GC} \simeq 0.4d$ , and  $\lambda_D \simeq 0.25d$ .

### 6.6.3 DEBYE–HÜCKEL REGIME

Broadly speaking (see Section 6.2.1), the PB equation can be linearized when the surface potential is small,  $|\Psi_s| \ll 1$ . In this case, the potential is small everywhere because it is a monotonous function that vanishes in the bulk. The DH solution for monovalent electrolytes has the general form

$$\Psi(z) = A \cosh(\kappa_D z) + B \sinh(\kappa_D z), \quad (6.81)$$

and the boundary conditions of Eq. (6.57) dictate the specific solution

$$\Psi(z) = -\frac{2}{\kappa_D \ell_{GC}} \frac{\cosh(\kappa_D z)}{\sinh(\kappa_D d/2)}. \quad (6.82)$$

In the DH regime, the potential is small, and the disjoining pressure (Eq. 6.58) can be expanded to second order in  $\Psi_m$ . As the first order vanishes from electroneutrality, we obtain

$$\Pi \simeq k_B T n_b \Psi_m^2 = \frac{k_B T}{2\pi \ell_B \ell_{GC}^2} \frac{1}{\sinh^2(\kappa_D d/2)}. \quad (6.83)$$

The DH regime can be further divided into two subclasses: DH<sub>1</sub> and DH<sub>2</sub>. For large separations,  $d \gg \lambda_D$ , the above expression reduces to

$$\Pi \simeq \frac{2k_B T}{\pi \ell_B \ell_{GC}^2} e^{-\kappa_D d}. \quad (6.84)$$

This DH<sub>1</sub> subregime is valid for  $d \gg \lambda_D$  and  $\ell_{GC} \gg \lambda_D$ .

In the other limit of small separations, the pressure is approximated by

$$\Pi \simeq \frac{2k_B T}{\pi \ell_B (\kappa_D \ell_{GC})^2} \frac{1}{d^2}. \quad (6.85)$$

The limits of validity for this DH<sub>2</sub> subregime are  $d \ll \lambda_D$  and  $\ell_{GC} \gg \lambda_D^2/d$  (see Table 6.2).

### 6.6.4 INTERMEDIATE REGIME

When  $d$  is the largest length scale in the system,  $d \gg \lambda_D$  and  $d \gg \ell_{GC}$ , the interaction between the membranes is weak, and one can use the superposition principle. This defines the *distal* region, where the midplane potential is obtained by adding the contributions from two identical charged single surfaces, located at  $z = \pm d/2$ .

In the distal region, the midplane potential is obtained from Eq. (6.33) by the above-mentioned superposition:

$$\Psi_m = -8\gamma e^{-\kappa_D d/2}. \quad (6.86)$$

Since  $\Psi_m$  is small, the pressure expression (Eq. 6.58) can be expanded to second order in  $\Psi_m$ , as was done in Eq. (6.83), giving

$$\Pi \simeq k_B T n_b \Psi_m^2 = 64 k_B T \gamma^2 n_b e^{-\kappa_D d}. \quad (6.87)$$

This osmotic pressure expression is valid for large distances,  $d \gg \lambda_D$  and  $d \gg \ell_{GC}$ , and partially holds for the DH<sub>1</sub> regime.

The *intermediate* regime is obtained by further assuming strongly charged surfaces,  $\lambda_D \gg \ell_{GC}$ . In this limit,  $\gamma = \tanh(-\Psi_s/4) \simeq 1$ , and the osmotic pressure is written as

$$\Pi \simeq \frac{8k_B T \kappa_D^2}{\pi \ell_B} e^{-\kappa_D d}. \quad (6.88)$$

**TABLE 6.2**  
**The Five Pressure Regimes of the Symmetric Two-Membrane System**

Pressure Regime	Pressure ( $\Pi$ )	Range of Validity
Ideal-Gas	$\frac{k_B T}{\pi \ell_B \ell_{GC}} \frac{1}{d}$	$\lambda_D/d \gg \ell_{GC}/\lambda_D \gg d/\lambda_D$
Gouy–Chapman	$\frac{\pi k_B T}{2 \ell_B} \frac{1}{d^2}$	$1 \gg d/\lambda_D \gg \ell_{GC}/\lambda_D$
Intermediate	$\frac{8 k_B T}{\pi \ell_B \lambda_D^2} e^{-d/\lambda_D}$	$d/\lambda_D \gg 1 \gg \ell_{GC}/\lambda_D$
Debye–Hückel (DH <sub>1</sub> )	$\frac{2 k_B T}{\pi \ell_B \ell_{GC}^2} e^{-d/\lambda_D}$	$d/\lambda_D \gg 1; \ell_{GC}/\lambda_D \gg 1$
Debye–Hückel (DH <sub>2</sub> )	$\frac{2 k_B T \lambda_D^2}{\pi \ell_B \ell_{GC}^2} \frac{1}{d^2}$	$1 \gg d/\lambda_D \gg \lambda_D/\ell_{GC}$

The intermediate regime is valid for  $d \gg \lambda_D \gg \ell_{GC}$  (see Table 6.2).

### 6.6.5 OTHER PRESSURE REGIMES

The pressure expression can be derived analytically in two other limits, which represent the two regimes obtained for the counterions-only case: the Ideal-Gas (IG) regime (Eq. 6.67),

$$\Pi \approx \frac{k_B T}{\pi \ell_B \ell_{GC}} \frac{1}{d}, \quad (6.89)$$

valid for  $\lambda_D^2/d \gg \ell_{GC} \gg d$ , and the Gouy–Chapman (GC) regime, (Eq. 6.68),

$$\Pi \approx \frac{\pi k_B T}{2 \ell_B} \frac{1}{d^2}, \quad (6.90)$$

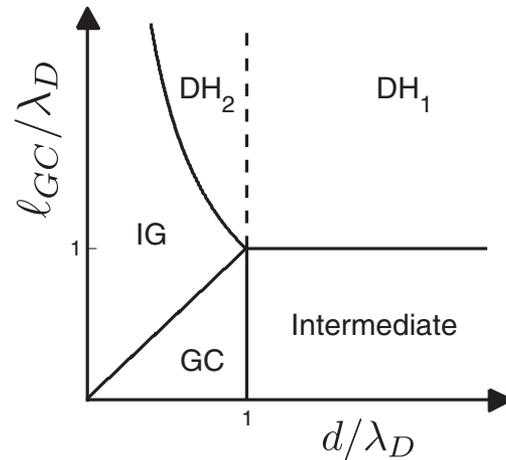
whose range of validity is  $\lambda_D \gg d \gg \ell_{GC}$ .

The five pressure regimes complete the discussion of the various limits as function of the two ratios:  $\ell_{GC}/\lambda_D$  and  $d/\lambda_D$ . They are summarized in Table 6.2 and plotted in Figure 6.10.

### 6.7 TWO ASYMMETRIC MEMBRANES, $\sigma_1 \neq \sigma_2$

For asymmetrically charged membranes,  $\sigma_1 \neq \sigma_2$ , the interacting membranes impose different boundary conditions. Such a system can model, for example, two surfaces that are coated with two different polyelectrolytes or two lipid membranes with different charge/neutral lipid compositions.

It is possible to have an overall attractive interaction between two asymmetric membranes, unlike the symmetric  $\sigma_1 = \sigma_2$  case. When  $\sigma_1$  and  $\sigma_2$  have the same sign, the boundary condition of Eq. (6.51) implies that  $\Psi'(d/2)$  has the opposite sign of  $\Psi'(-d/2)$ . Since  $\Psi'$  is monotonous, it means that there is a point between the plates for which  $\Psi' = 0$ . The osmotic pressure,  $\Pi$  of Eq. (6.55), calculated at this special point, has only an entropic contribution and is

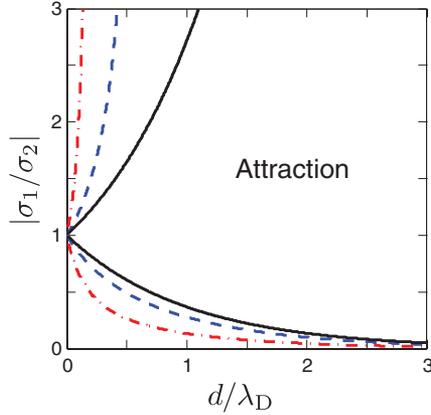


**FIGURE 6.10** Schematic representation of the various regimes of the PB equation for two flat and equally charged membranes at separation  $d$ . We plot the four different pressure regimes: Ideal-Gas (IG), Gouy–Chapman (GC), intermediate, and Debye–Hückel (DH). The two independent variables are the dimensionless ratios  $d/\lambda_D$  and  $\ell_{GC}/\lambda_D$ . The four regimes are detailed in Table 6.2. The DH regime is further divided into two subregimes: DH<sub>1</sub> for large  $d/\lambda_D$  and DH<sub>2</sub> for small  $d/\lambda_D$ .

positive for any intermembrane separation,  $d$ , just as in the  $\sigma_1 = \sigma_2$  case.

However, when  $\sigma_1$  and  $\sigma_2$  have opposite signs,  $\Psi'$  is always negative between the two membranes, and the sign of the pressure can be either positive (repulsive) or negative (attractive). A crossover between repulsive and attractive pressure occurs when  $\Pi = 0$  and depends on four system parameters:  $\sigma_{1,2}$ ,  $\lambda_D$ , and  $d$  (see Figure 6.11).

Although the general expression for  $\Pi(d)$  cannot be cast in an analytical form, a closed-form criterion exists for the crossover pressure,  $\Pi = 0$ , for any amount of salt (Ben-Yaakov et al. 2007). The crossover criterion has two rather simple limits: for the linearized DH (high salt) limit, the general criterion reduces to the well-known result of Parsegian and Gingell (1972), while in the counterion-only



**FIGURE 6.11** Crossover from attraction to repulsion,  $\Pi(d)=0$ , for two oppositely charged membrane,  $\sigma_1 \cdot \sigma_2 < 0$ , in the  $(|\sigma_1/\sigma_2|, d/\lambda_D)$  plane. The solid lines show the crossover in the high-salt DH limit, Eq. (6.95). For lower salinity, derived from the full PB theory, Eq. (6.107), the attractive region is increased as is seen in the examples we choose:  $\ell_2/\lambda_D = 0.2$  (red dash-dotted line) and  $\ell_2/\lambda_D = 0.7$  (blue dashed line).

limit, another analytical expression has been derived more recently by Lau and Pincus (1999).

### 6.7.1 THE DEBYE–HÜCKEL REGIME

The crossover criterion between attraction and repulsion has an analytical limit for high salinity, (Parsegian and Gingell 1972). We repeat here the well-known argument (Ben-Yaakov and Andelman 2010) where the starting point is the linear DH limit (Eq. 6.14) of the full PB equation.

The DH equation for monovalent electrolyte in planar geometries has a solution (Eq. 6.81), for which the boundary conditions of Eq. (6.51) yield

$$\Psi(z) = \frac{2\pi\ell_B}{\kappa_D e} \left[ \frac{\sigma_1 + \sigma_2}{\sinh(\kappa_D d/2)} \cosh(\kappa_D z) + \frac{\sigma_2 - \sigma_1}{\cosh(\kappa_D d/2)} \sinh(\kappa_D z) \right]. \quad (6.91)$$

In the DH regime, the pressure expression (Eq. 6.55) can be expanded in powers of the electrostatic potential,  $\Psi$ . Keeping only terms of order  $\Psi^2$ , the pressure can be written as

$$\Pi \approx \frac{k_B T}{2\pi\ell_B \sinh^2(\kappa_D d)} \left[ \frac{1}{\ell_1^2} + \frac{1}{\ell_2^2} \pm \frac{2}{\ell_1 \ell_2} \cosh(\kappa_D d) \right], \quad (6.92)$$

where  $\ell_{1,2} = e/(\kappa_D |\sigma_{1,2}|)$  are the two Gouy–Chapman lengths corresponding to the two membranes with  $\sigma_1$  and  $\sigma_2$ , respectively. The  $\pm$  sign of the last term corresponds to the two situations:  $\sigma_1 \cdot \sigma_2 > 0$  and  $\sigma_1 \cdot \sigma_2 < 0$ , respectively.

This  $\Pi$  expression can be simplified in two limits. For small separation,  $d \ll \lambda_D$ , the expansion of the hyperbolic functions yields a power law divergence  $\sim d^{-2}$  for  $d \rightarrow 0$ ,

$$\Pi \approx \frac{k_B T}{2\pi\ell_B} \left[ \left( \frac{1}{\kappa_D \ell_1} \pm \frac{1}{\kappa_D \ell_2} \right)^2 \frac{1}{d^2} \pm \frac{1}{\ell_1 \ell_2} \right] > 0. \quad (6.93)$$

Clearly, it is positive definite (hence repulsive) for both  $\pm$  signs. However, when  $\sigma_1 = -\sigma_2$  (the antisymmetric case with  $\sigma_1 \cdot \sigma_2 < 0$ ), the pressure goes to a negative constant (independent of  $d$ ),  $\Pi = -k_B T / (2\pi\ell_B \ell_1 \ell_2)$ .

For the opposite limit of large separation  $d \gg \lambda_D$ , the pressure decays exponentially, while its sign depends on the sign of  $\sigma_1 \cdot \sigma_2$ ,

$$\Pi \approx \pm \frac{k_B T}{\pi\ell_B \ell_1 \ell_2} e^{-\kappa_D d}, \quad (6.94)$$

and thus, it is attractive for  $\sigma_1 \cdot \sigma_2 < 0$ .

The attraction/repulsion crossover is calculated from the zero-pressure condition of Eq. (6.92), while keeping in mind that attraction is possible only for oppositely charged membranes,  $\sigma_1 \cdot \sigma_2 < 0$  (see the beginning of this section):

$$e^{-\kappa_D d} < \left| \frac{\sigma_1}{\sigma_2} \right| < e^{\kappa_D d}. \quad (6.95)$$

This is exactly the result obtained by Parsegian and Gingell (1972). Interestingly, in the linear DH case, the crossover depends only on the ratio of the two surface charges  $|\sigma_1/\sigma_2|$  and not on their separate values, as can be seen in Figure 6.11. For comparison, we plot (with dashed and dash-dotted lines on the same figure) two examples of low-salt crossovers, as calculated from the general criterion presented below for the full PB theory (Section 6.7.4). The low-salt line has a smaller repulsive region. Increasing the salt concentration increases the repulsion region due to screening of electrostatic interactions. The repulsive region increases till it reaches the Parsegian–Gingell result for the DH limit (solid line in Figure 6.11).

### 6.7.2 THE DEBYE–HÜCKEL REGIME WITH CONSTANT SURFACE POTENTIAL

So far we have solved the PB equation using the constant charge boundary conditions. However, constant potential boundary conditions are appropriate when the surfaces are metal electrodes, and it is important to understand this case as well.

Let us examine the effect of constant surface potential on the pressure. For simplicity, we will focus on the linearized PB equation (DH) in the asymmetric membrane case for monovalent electrolyte. We still refer to the set-up as in Figure 6.7. The two membranes at  $z = \mp d/2$  are held at different values of constant surface potential,  $\Psi_{1,2}$ :

$$\begin{aligned} \Psi|_{z=-d/2} &= \Psi_1, \\ \Psi|_{z=d/2} &= \Psi_2. \end{aligned} \quad (6.96)$$

Applying the DH solution of Eq. (6.81) with the boundary conditions of Eq. (6.96) leads to

$$\Psi(z) = \frac{\Psi_1 + \Psi_2}{2\cosh(\kappa_D d/2)} \cosh(\kappa_D z) + \frac{\Psi_1 - \Psi_2}{2\sinh(\kappa_D d/2)} \sinh(\kappa_D z), \quad (6.97)$$

and expanding the pressure  $\Pi$  from Eq. (6.56) to the second order in powers of  $\Psi_{1,2}$  yields

$$\Pi \approx \frac{k_B T n_b}{\sinh^2(\kappa_D d)} (2\Psi_2 \Psi_1 \cosh(\kappa_D d) - \Psi_2^2 - \Psi_1^2). \quad (6.98)$$

This expression is similar to the one obtained for constant surface charge (Eq. 6.92). Indeed, for large separations,  $d \gg \lambda_D$ , the relative sign of  $\Psi_2$  and  $\Psi_1$  determines the sign of the pressure

$$\Pi \approx 2k_B T n_b \Psi_2 \Psi_1 e^{-\kappa_D d}, \quad (6.99)$$

as for the large-separation behaviour of the constant-charge case.

However, for small separations,  $d \ll \lambda_D$ , the pressure is different than for the constant surface-charge case,

$$\Pi \approx -\frac{k_B T (\Psi_2 - \Psi_1)^2}{8\pi\ell_B} \frac{1}{d^2} + k_B T n_b \Psi_2 \Psi_1. \quad (6.100)$$

It yields a pure attractive (negative) pressure that diverges as  $\sim 1/d^2$ , and does not depend on  $n_b$ . For the special symmetric case  $\Psi_2 = \Psi_1$ , at those small  $d$ , the pressure does not diverge and reaches a positive constant,  $\Pi > 0$ , proportional to  $n_b$ . Unlike the constant-charge case, here the counterion concentration remains constant near each of the membranes, because it depends only on the surface potential through the Boltzmann factor (Ben-Yaakov and Andelman 2010). However, the induced surface charge ( $\sigma \propto \Psi'_s$ , Eq. (6.97)) diverges when the membranes are brought closer together, resulting in a diverging electrostatic attraction.

Note that the crossover from repulsive to attractive pressure is obtained for zero pressure in Eq. (6.98) and is possible only for potentials of the same sign,  $\Psi_2 \cdot \Psi_1 > 0$  including  $\Psi_1 = \Psi_2$ . The condition for attraction reads

$$e^{-\kappa_D d} < \frac{\Psi_2}{\Psi_1} < e^{\kappa_D d}. \quad (6.101)$$

For potentials of opposite sign,  $\Psi_2 \cdot \Psi_1 < 0$ , the pressure is purely attractive.

### 6.7.3 COUNTERIONS ONLY

In the absence of an external salt reservoir, the only mobile ions in the solution are monovalent counterions with density  $n(z)$ , such that the system is charge neutral,

$$\sigma_1 + \sigma_2 = -e \int_{-d/2}^{d/2} n(z) dz. \quad (6.102)$$

For the assumed overall negative charge on the two membranes,  $\sigma_1 + \sigma_2 < 0$ , the counterions are positive,  $z_+ = 1$ .

The PB equation for the two-membrane system is the same as for the single membrane (Eq. 6.18), with the boundary condition as in Eq. (6.51). The osmotic pressure (Eq. 6.55) reduces here to

$$\Pi = -\frac{k_B T}{8\pi\ell_B} (\Psi')^2 + k_B T n_0 e^{-\Psi(z)}, \quad (6.103)$$

where  $n_0$  is defined as the reference density for which  $\Psi = 0$ . This equation is a first-order ordinary differential equation and can be integrated. Nevertheless, its solution depends on the sign of the osmotic pressure. We will not present here the solution of the PB equation but rather discuss the crossover between attractive and repulsive pressures. This crossover is obtained by solving Eq. (6.103) with  $\Pi = 0$ . As the total surface charge is chosen to be negative,  $\sigma_1 + \sigma_2 \leq 0$ , and attraction occurs only for  $\sigma_1 \cdot \sigma_2 < 0$ , we choose  $\sigma_1$  to be negative and  $\sigma_2$  to be positive.

Integrating this equation and using the boundary condition at  $z = -d/2$ , we obtain the same electrostatic potential profile as in the single membrane case (Eq. 6.22), with a shifted  $z$ -axis origin:  $z \rightarrow z + d/2$  and  $\ell_{GC} \rightarrow \ell_1$  (with  $\ell_1$  defined as before):

$$\Psi = \Psi_0 + 2 \ln(z + \ell_1 + d/2). \quad (6.104)$$

The second boundary condition at  $z = d/2$  gives a relation between  $d$  and  $\sigma_{1,2}$ . The condition for attraction can be expressed in terms of the surface densities (Kanduč et al. 2008):

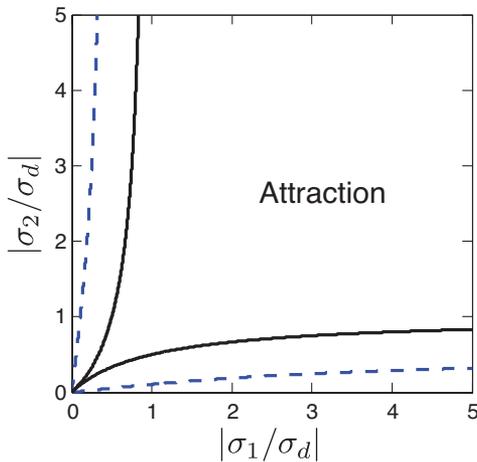
$$|\sigma_1| - |\sigma_2| < \frac{|\sigma_2 \sigma_1|}{\sigma_d}, \quad (6.105)$$

where  $\sigma_d \equiv e/(2\pi\ell_B d)$ .

The crossover between attraction and repulsion is plotted in Figure 6.12. Two crossover lines separate the central attraction region from two repulsion ones. The upper one lies above the diagonal,  $|\sigma_1| > |\sigma_2|$ , and corresponds directly to the condition of Eq. (6.105). A second crossover line lies in the lower wedge below the diagonal,  $|\sigma_1| < |\sigma_2|$ . It corresponds to the crossover of the complementary problem of an overall positive surface charge,  $\sigma_1 + \sigma_2 > 0$ , and negative counterions. Note that the figure is symmetric about the principal diagonal,  $|\sigma_1| \leftrightarrow |\sigma_2|$ , as expected. The condition of attraction, irrespectively of the sign of  $\sigma_1 + \sigma_2$ , can be written as

$$\left| |\sigma_1| - |\sigma_2| \right| < \frac{|\sigma_1 \sigma_2|}{\sigma_d}, \quad (6.106)$$

as was obtained by Lau and Pincus (1999).



**FIGURE 6.12** Regions of attraction ( $\Pi < 0$ ) and repulsion ( $\Pi > 0$ ) for the counterion-only case, plotted in terms of the two rescaled charge densities,  $|\sigma_1/\sigma_d|$  and  $|\sigma_2/\sigma_d|$ , where  $\sigma_d = e/2\pi\ell_B d$ . The figure is plotted for oppositely charged membranes,  $\sigma_1 \cdot \sigma_2 < 0$ , and is symmetric about the diagonal  $|\sigma_1| = |\sigma_2|$ . The two solid lines delimit the boundary between repulsion and attraction in the no-salt limit,  $n_b \rightarrow 0$ , Eq. (6.106). For comparison, we also plot the crossover between attraction and repulsion for finite  $n_b$  from Eq. (6.107) for  $d/\lambda_D = 2.2$  (blue dashed line). For the case of  $\sigma_1 \cdot \sigma_2 > 0$ ,  $\Pi$  is always repulsive and there is no crossover.

#### 6.7.4 ATTRACTION/REPULSION CROSSOVER

We now calculate the general criterion of the attractive-to-repulsive crossover. The  $\Pi = 0$  pressure between two membranes located at  $z = \pm d/2$  can be mapped exactly into the problem of a single membrane at  $z = 0$  in contact with the same electrolyte reservoir. The only difference is that beside the boundary at  $z = 0$ , there is another boundary at  $z = d$ . The mapping to the single-membrane case is possible as the osmotic pressure of a single membrane is zero. This equivalence can be checked by substituting  $\Pi = 0$  in Eq. (6.55) to recover the first integration of the PB equation for the single membrane system with added 1:1 electrolyte (Eq. 6.26).

The potential can be written as in Eq. (6.27) with  $\gamma \rightarrow \gamma_1 = \sqrt{1 + (\kappa_D \ell_1)^2} - \kappa_D \ell_1$ . This solution already satisfies the boundary condition,  $\Psi'(0) = -4\pi\ell_B \sigma_1/e$ , but another boundary condition at  $z = d$  needs to be satisfied as well,  $\Psi'(d) = 4\pi\ell_B \sigma_2/e$ .

Attraction will occur only for charged membranes of different sign,  $\sigma_1 \cdot \sigma_2 < 0$ . Using the boundary condition at  $z = d$  for the two cases,  $\sigma_1 < 0$  and  $\sigma_1 > 0$ , determines the region of attraction,  $\Pi < 0$ , by the inequalities (Ben-Yaakov et al. 2007)

$$e^{-\kappa_D d} < \frac{\gamma_2}{\gamma_1} < e^{\kappa_D d}, \quad (6.107)$$

with  $\gamma_2 = \sqrt{1 + (\kappa_D \ell_2)^2} - \kappa_D \ell_2$ . It can be shown that the above general expression (Eq. 6.107) reduces to the

expression of Eq. (6.106), in the limit of counterion only and to that of Eq. (6.95) in the high-salt (DH) limit.

A similar general crossover criterion can also be obtained for constant potential boundary conditions. As we explained above, the crossover condition maps to the single membrane problem, yielding the generalized relation of Eq. (6.30),  $\gamma_{1,2} = \pm \tanh(\Psi_{1,2}/4)$ . The  $\pm$  sign is chosen such that  $\gamma_{1,2}$  is positive. For opposite surface potentials,  $\Psi_1 \cdot \Psi_2 < 0$ , there is a point between the membranes in which the potential vanishes. The osmotic pressure of Eq. (6.55), calculated at this point, has only the negative Maxwell stress contribution, and therefore, it is always *attractive*. On the other hand, for  $\Psi_1 \cdot \Psi_2 > 0$ , the following condition on  $\Psi_1$  and  $\Psi_2$  results in an attraction:

$$e^{-\kappa_D d} < \frac{\tanh(\Psi_2/4)}{\tanh(\Psi_1/4)} < e^{\kappa_D d}. \quad (6.108)$$

In Figure 6.13, we show the osmotic pressure,  $\Pi$ , in units of  $k_B T / (4\pi\ell_B \lambda_D^2)$ , as a function of the (dimensionless) intermembrane separation,  $d/\lambda_D$ . The pressure is calculated for several values of constant charge and constant potential boundary conditions. Three types of pressure profiles are seen in the figure: attractive, repulsive, and the crossover between attraction and repulsion.

#### 6.8 CHARGE REGULATION

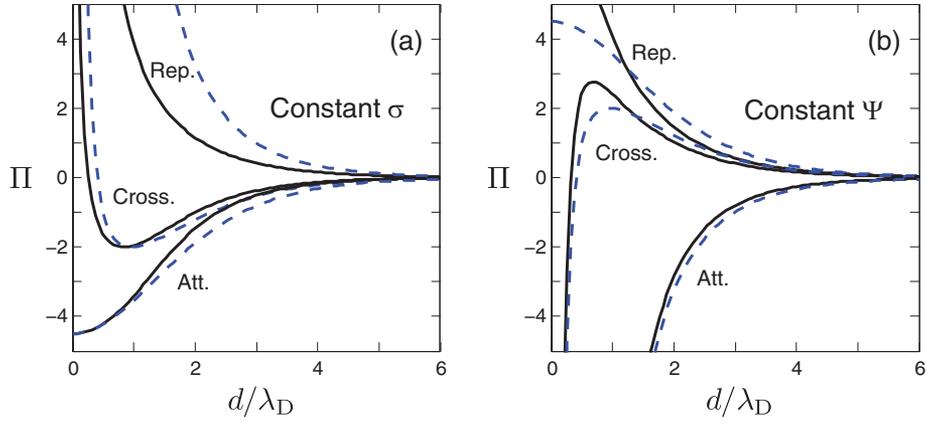
As discussed in Section 6.7.2, the difference between constant surface potential,  $\Psi_s$ , and constant surface charge density,  $\sigma$ , is large when the distance between the two membranes is of order of the Debye screening length,  $\lambda_D$ , or smaller. Ninham and Parsegian (1971) considered an interesting intermediate case of great practical importance. Membranes with ionizable groups that can release ions into the aqueous solution or trap them – a situation intermediate between a constant  $\sigma$ , describing inert ionic groups on the membrane, and constant  $\Psi_s$ , relevant for a surface (an electrode or a membrane) held at a constant potential by an external potential source.

Let us consider a system where the membrane is composed of ionizable groups (lipids), and each of them can release a counterion into the solution. This surface dissociation/association (see Figure 6.14) is described by the reaction:

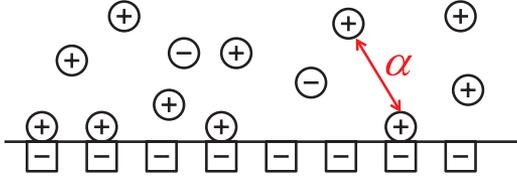


where A denotes a surface site that can be either ionized ( $A^+$ ) or neutral (AB). The process of membrane association/dissociation is characterized by a kinetic constant  $K_d$  through the law of mass action:

$$K_d = \frac{[A^+][B^-]_s}{[AB]}, \quad (6.110)$$



**FIGURE 6.13** Osmotic pressure,  $\Pi$ , in units of  $k_B T / (4\pi\ell_B\lambda_D^2)$ , as function of the dimensionless intermembrane separation,  $d/\lambda_D$ . We present the solution of the nonlinear (solid black lines) PB equation and the linear (blue dashed lines) DH equation for two boundary conditions: (a) constant surface charge, and (b) constant surface potential. In each of the figure parts we show three profiles: repulsive, crossover, and attractive. In (a), the boundary conditions for the repulsive, crossover, and attractive profiles are  $\sigma_1 = \sigma_2 = 3$ ;  $\sigma_1 = 3$  and  $\sigma_2 = -2$ ; and  $\sigma_1 = 3$  and  $\sigma_2 = -3$ , respectively, where  $\sigma$  is given in units of  $e / (4\pi\ell_B\lambda_D)$ . In (b), the boundary conditions for the repulsive, crossover, and attractive profiles are  $\Psi_1 = \Psi_2 = 3$ ;  $\Psi_1 = 3$  and  $\Psi_2 = 2$ ; and  $\Psi_1 = 3$  and  $\Psi_2 = -3$ , respectively. Note that the nonlinear PB solution of  $\Pi$  for the symmetric (repulsive) osmotic pressure for constant potential reaches a constant value as  $d \rightarrow 0$ , like in the DH case, but with a different value,  $\Pi(d \rightarrow 0) \simeq 9.07$  in units of  $k_B T / (4\pi\ell_B\lambda_D^2)$  (not shown in the figure).



**FIGURE 6.14** Illustration of the charge regulation boundary condition for cation association/dissociation energy  $\alpha$ .

where  $[A^+]$ ,  $[B^-]_s$  and  $[AB]$  denote the three corresponding surface concentrations (per unit volume). We define  $\phi_s$  to be the area fraction of the  $A^+$  ions, related to  $\sigma > 0$ , the membrane charge density by  $\phi_s = \sigma a^2 / e \sim [A^+]$  and  $1 - \phi_s \sim [AB]$ , where  $a^2$  is the surface area per charge. The equilibrium condition of Eq. (6.110) is then written as

$$K_d = \frac{\phi_s}{1 - \phi_s} [B^-]_s. \quad (6.111)$$

As the counterions are released into the ionic solution, the relation between surface and bulk  $B^-$  concentrations,  $[B^-]_s$  and  $[B^-]_\infty$ , is obtained via the Boltzmann distribution,  $[B^-]_s = [B^-]_\infty \exp(\Psi_s)$  where  $[B^-]_\infty = n_b$  is the bulk salt concentration. We note that in this section we choose  $\sigma > 0$  and it implies  $\Psi_s > 0$ , and the reference potential in the bulk is set to be  $\Psi = 0$ . The concentration of dissociated  $B^-$  ions at the surface is  $[B^-]_s$  and it is equal to  $n_s^- = n_-(z \rightarrow 0)$ . Note that this is *not* the charge concentration of the membrane itself (which comprises the bound ionic groups and is proportional to  $\phi_s$ ) but is the concentration of mobile ions evaluated at the submembrane position (just as in Section 6.3.3), where the standard PB equation holds. One can then write the area fraction of the membrane ionized sites as

$$\phi_s = \frac{K_d}{K_d + n_b e^{\Psi_s}}. \quad (6.112)$$

It is useful to introduce a surface interaction parameter  $\alpha = \ln(a^3 K_d)$  instead of using the kinetic constant (see Figure 6.14). This gives the adsorption isotherm for  $\phi_s$  (the fraction of  $A^+$  groups on the membrane):

$$\phi_s = \frac{1}{1 + \phi_b e^{-\alpha + \Psi_s}}, \quad (6.113)$$

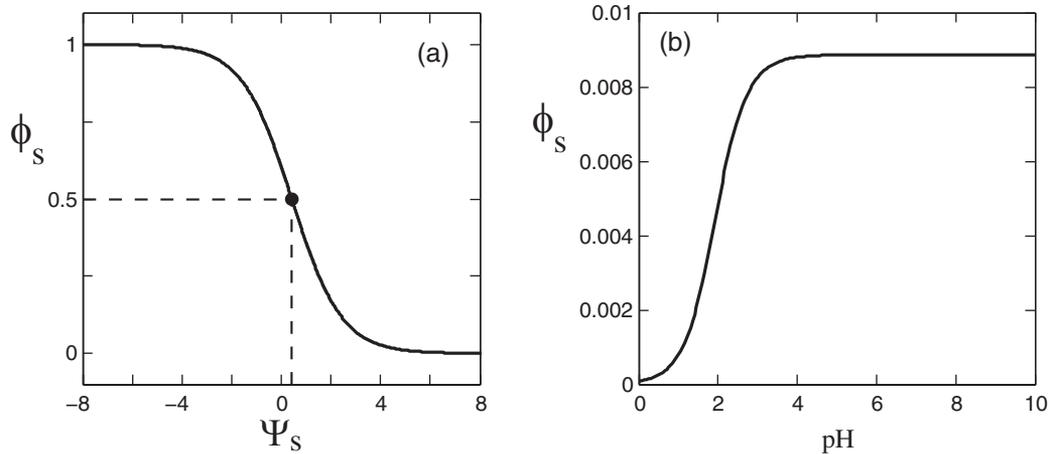
where  $\phi_b = a^3 n_b$  as before. The typical shape of  $\phi_s$  as function of  $\Psi_s$  is a sigmoid and is shown in Figure 6.15a for  $a = 0.3$  nm,  $n_b = 0.1$  M, and  $\alpha = -6$  (pK  $\simeq 0.82$ ). The fraction of charge groups,  $\phi_s$ , varies between  $\phi_s = 1$  (fully charged) and  $\phi_s = 0$  (neutral) as  $\Psi_s$  varies from negative values to positive ones. For the half-filled surface charge,  $\phi_s = 0.5$ , the surface potential is  $\Psi_s^* = \alpha - \ln \phi_b$ . At this special point, the slope of  $\phi_s(\Psi_s)$  is exactly  $-0.25$ . The differential capacitance as discussed in Section 6.4 can be cast into a simple form for the charge regulation case

$$C_{CR} = \frac{d\sigma}{d\Psi_s} = \frac{e^2}{k_B T a^2} \frac{d\phi_s}{d\Psi_s} = \frac{e^2}{k_B T a^2} \phi_s (1 - \phi_s). \quad (6.114)$$

It follows that for the sigmoid shape of  $\phi_s(\Psi_s)$ , the differential capacitance has a unimodal shape with a universal maximum whose value is  $e^2 / (4k_B T a^2)$ .

Similarly, for the fraction of neutral AB groups on the membrane,  $1 - \phi_s$ , we write

$$1 - \phi_s = \frac{\phi_b}{\phi_b + e^{\alpha - \Psi_s}}. \quad (6.115)$$



**FIGURE 6.15** In (a), the area fraction of charge groups on the membrane,  $\phi_s$ , is plotted as a function of the dimensionless electrostatic potential,  $\Psi_s$ , using Eq. (6.113). The parameters used are  $a=0.3$  nm,  $n_b=0.1$  M, and  $\alpha=-6$  ( $\text{pK} \approx 0.82$ ). The symmetric point of  $\phi_s=0.5$  occurs at  $\Psi_s^* \approx 0.425$ , and the slope there is exactly  $-0.25$ . In (b), we present the area fraction of charge groups on the membrane,  $\phi_s$ , as a function of pH for surface binding  $A^- + H^+ \rightleftharpoons AH$  (see “Surface pH and pK” box below). The parameters used are:  $a=0.5$  nm and  $\alpha=-12$  ( $\text{pK} \approx 4.1$ ).

Equation (6.115) is the *Langmuir–Davies isotherm* (Davies 1958) and is an extension of the Langmuir adsorption isotherm (see, e.g., Adamson and Gast 1997) for charged adsorbing particles. This can be understood because  $1 - \phi_s$  is the fraction of the membrane AB neutral groups. Hence, it effectively describes the adsorption of  $B^-$  ions onto a charged membrane.

### SURFACE pH AND pK

When an acidic reservoir exchanges  $H^+$  ions with the membrane, the membrane chemical reaction is  $A^- + H^+ \rightleftharpoons AH$ , and the same local equilibrium of Eq. (6.110) can be expressed in terms of three logarithms:  $\text{pK} \equiv -\log_{10} K_d = -\log_{10}(a^{-3} e^\alpha)$  where the close packing density  $a^{-3}$  is measured in molar,  $\text{pH}_s \equiv -\log_{10}[H^+]_s$ , and  $\text{pH} \equiv -\log_{10}[H^+]_\infty$ . The regular pH measures the acidic strength in the reservoir, while the membrane pK (or  $K_d$ ) is a fixed (and usually unknown) parameter that depends on the membrane as well as on the binding  $H^+$  ions. Equation (6.111) can then be written as

$$\frac{\phi_s}{1 - \phi_s} = 10^{-\text{pK} + \text{pH}} e^{\Psi_s}. \quad (6.116)$$

When half of the membrane is charged,  $\phi_s = 0.5$ , the given pK relates the solution pH with the surface potential:

$$\Psi_s = (\text{pH} - \text{pK}) \ln(10). \quad (6.117)$$

In Figure 6.15b, we show the dependence of  $\phi_s$  on pH for process of protonation/deprotonation for the case of one membrane with:  $a = 0.5$  nm and  $\alpha = -12$  ( $\text{pK} \approx 4.1$ ). The typical shape of  $\phi_s(\text{pH})$  is a sigmoid and is obtained by solving numerically Eq. (6.118) with the charge regulation boundary condition of Eq. (6.113).

The Langmuir–Davies isotherm (Eq. 6.115) relates the self-adjusting surface charge fraction  $\phi_s$  and surface potential  $\Psi_s$  with  $K_d$  (or  $\alpha$ ) and the bulk density,  $n_b$ . In order to find  $\Psi_s$  and  $\phi_s$  separately as function of  $K_d$  and  $n_b$ , one needs to find the electrostatic relation between them. This relation can be obtained from the Grahame equation (6.36), introduced in Section 6.3.3:

$$w\phi_s = \sinh(\Psi_s/2); \quad w \equiv \frac{2\pi\ell_B\lambda_D}{a^2}, \quad (6.118)$$

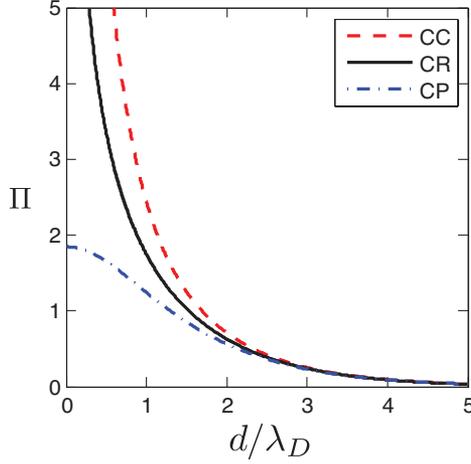
where we expressed the Grahame equation as  $\phi_s = \phi_s(\Psi_s)$  in terms of the Bjerrum and Debye lengths via a dimensionless parameter,  $w$ . By inverting the relation,  $\Psi_s = \Psi_s(\phi_s)$ , we get

$$e^{\Psi_s/2} = w\phi_s + \sqrt{1 + (w\phi_s)^2}. \quad (6.119)$$

Recall that this Grahame equation (6.36) is obtained from the PB equation for one planar membrane and gives a relationship between the density  $\sigma$  at the membrane and  $n_s^\pm$  at the subsurface layer. The Grahame equation depends only on the electrostatics properties and applies for any boundary condition: constant  $\sigma$ , constant  $\Psi_s$ , or the present case of charge regulation due to association/dissociation of the ionizable groups on the membrane. These two equations (Eqs. 6.115 and 6.118) determine completely the surface potential,  $\Psi_s = \Psi_s(K_d, n_b)$ , and the surface charge fraction,  $\phi_s = \phi_s(K_d, n_b)$ .

It is instructive to take the two limits of large and small  $\phi_s$ . When the ionizable surface sites are almost fully dissociated,  $\phi_s \approx 1$ ,  $K_d$  is large enough so that  $K_d \gg n_b \exp(\Psi_s)$ . In this limit, Eqs. (6.112) and (6.119) reduce to

$$\begin{aligned} \phi_s &\approx 1 - \frac{n_b}{K_d} e^{\Psi_s}, \\ \Psi_s &\approx 2 \ln\left(w + \sqrt{1 + w^2}\right), \end{aligned} \quad (6.120)$$



**FIGURE 6.16** The pressure for two symmetric membranes in units of  $k_B T / (4\pi\ell_B \lambda_D^2)$ , in the presence of three different boundary conditions: constant charge (CC), constant potential (CP), and charge regulation (CR). The pressure inequality, seen in the figure,  $\Pi_{CC} > \Pi_{CR} > \Pi_{CP}$ , is an inequality that holds in general. The parameters used are  $a = 0.5$  nm,  $n_b = 0.1$  M, and  $\alpha = -6$  (pK  $\approx 1.48$ ).

and for the additional requirement of large  $w$  (or  $\lambda_D \gg a^2/\ell_B$ )

$$\Psi_s \approx 2 \ln(2w) = 2 \ln\left(\frac{8\pi\ell_B \lambda_D}{a^2}\right). \quad (6.121)$$

In the opposite limit, only a small fraction of the ionizable surface groups are dissociated,  $\phi_s \ll 1$ , which results in  $K_d \ll n_b \exp(\Psi_s)$  and

$$\phi_s \approx \frac{K_d}{n_b} e^{-\Psi_s} \ll 1. \quad (6.122)$$

If, in addition to small  $\phi_s$  also  $w\phi_s \ll 1$ , expanding the right-hand side of Eq. (6.119) gives

$$\Psi_s \approx 2w\phi_s \ll 1. \quad (6.123)$$

As stated in the beginning of this section, the charge regulation boundary condition lies between the constant charge and constant potential boundary conditions. In other words, the osmotic pressure obtained for constant charge (CC) boundary condition,  $\Pi_{CC}$ , will always be more repulsive than the pressure,  $\Pi_{CR}$ , of charge regulation (CR). Furthermore, the latter is more repulsive than  $\Pi_{CP}$ , the pressure for constant potential (CP), yielding  $\Pi_{CC} > \Pi_{CR} > \Pi_{CP}$ .

A calculated example manifests this fact and can be seen in Figure 6.16, where we show the osmotic pressure,  $\Pi$ , in units of  $k_B T / (4\pi\ell_B \lambda_D^2)$  as a function of the dimensionless inter-membrane separation,  $d/\lambda_D$ . We use Eqs. (6.69) and (6.113) to obtain numerically the surface potential,  $\Psi_s$ , for the symmetric charge regulation boundary condition for any intermembrane separation,  $d/\lambda_D$ . The midplane potential is obtained from Eq. (6.79), and the pressure is calculated via Eq. (6.58). The difference in the osmotic pressure for

the three boundary conditions arises only for small separations,  $d \approx \lambda_D$ , while in the large separation limit,  $d \gg \lambda_D$ , the three pressures coincide. From the extrapolation of the CR surface potential at large separations,  $\Psi_s(d \rightarrow \infty)$ , we find the constant surface potential to be,  $\Psi_s^{(\text{const})} \approx -1.7$ , while the constant surface charge is  $\sigma_{\text{const}} \approx -e / 4.35 \text{ nm}^2$ .

The behaviour of  $\Pi_{CP}$  for small distances,  $d \ll \lambda_D$ , is very different from the behaviour of  $\Pi_{CC}$  and of  $\Pi_{CR}$ , as  $\Pi_{CP}$  saturates at a value of  $\Pi_{CP} \approx 1.85$  (in units of  $k_B T / (4\pi\ell_B \lambda_D^2)$ ). For  $d/\lambda_D \rightarrow 0$  and finite  $\ell_{GC}$ ,  $\Pi_{CC}$  is always in the ideal gas regime (see Table 6.2), thus diverging as  $\Pi_{CC} \sim 1/d$ . Note that the osmotic pressure  $\Pi_{CR}$  also diverges when  $d/\lambda_D \rightarrow 0$ , but only as  $\Pi_{CR} \sim 1/d^{1/2}$ , which is weaker than for  $\Pi_{CC}$  (Markovich, Andelman and Podgornik 2016).

### 6.8.1 CHARGE REGULATION VIA FREE ENERGY

The Langmuir–Davies isotherm (Eq. 6.115) can also be derived from a free-energy minimization (Diamant and Andelman 1996). It is done by including a surface free energy,  $F_s$ , to account for the association/dissociation of ions onto/from the membrane. The total free energy is  $F_t = F_v + F_s$ , where the volume contribution,  $F_v$ , is the same as in Eq. (6.11) and the surface free energy is

$$a^2 F_s / (A k_B T) = \Psi_s \phi_s + \phi_s \ln \phi_s + (1 - \phi_s) \ln(1 - \phi_s) + \alpha(1 - \phi_s), \quad (6.124)$$

where  $A$  is the lateral membrane area and  $a$  is its thickness. The first term describes the coupling between the surface charge density  $\sigma = e\phi_s/a^2$  (taken as positive) and the surface potential  $\Psi_s = e\psi_s/k_B T$ . The second and third terms describe the mixing entropy of dissociated (charged) surface sites of fraction  $\phi_s$  and associated (thus neutral) ones of fraction  $1 - \phi_s$ . Finally, the fourth term is proportional to  $1 - \phi_s$ , the amount of  $B^-$  ions adsorbing onto the membrane. It accounts for the excess surface interaction as an ion binds onto the membrane creating a neutral AB group. The surface interaction parameter,  $\alpha$ , is the same as the one defined after Eq. (6.112).

In order to obtain the surface isotherm, one needs to take the variation of  $F_t$  with respect to  $\phi_s = \sigma a^2/e$ ,  $\delta F_t / \delta \phi_s = \delta F_s / \delta \phi_s = \mu_s$ , giving

$$\Psi_s + \ln\left(\frac{\phi_s}{1 - \phi_s}\right) + (\beta\mu_s - \alpha) = 0. \quad (6.125)$$

For a dilute ionic solution, the chemical potential of the ions is related to their bulk density by  $\mu_b = k_B T \ln \phi_b$ , as explained after Eq. (6.6). In thermodynamical equilibrium, the chemical potential is equal throughout the solution, hence,  $\mu_s = \mu_b$ , and by rearranging Eq. (6.125) the Langmuir–Davies isotherm emerges,

$$1 - \phi_s = \frac{\phi_b}{\phi_b + e^{-\Psi_s}}. \quad (6.126)$$

The above equation is exactly the Langmuir–Davies isotherm of Eq. (6.115).

The advantage of the free-energy formulation presented in this section over the chemical equilibrium one presented earlier is that the former can be generalized to other cases of surface interaction, such as cooperativity between the surface sites modelled by adding a  $b(1-\phi_s)^2$  term to  $F_s$ , adsorbing of several ion types with different ion–surface interactions,  $\alpha$ , and other extensions of the simple charge regulation mechanism (Diamant and Andelman 1996, Ariel, Diamant and Andelman 1999). Recently, the free-energy formulation was also used to study the collective behaviour of mobile charge-regulation macromolecules (Markovich, Andelman and Podgornik 2017b, Avni et al. 2018, Hallett et al. 2018, Avni, Andelman and Podgornik 2019, Avni, Podgornik and Andelman 2020).

## 6.9 VAN DER WAALS' INTERACTIONS

Long-range van der Waals' (vdW) interactions between molecules are universal and result from the molecular dipolar fluctuations (for details see Parsegian 2005, Bordag et al. 2009). These fluctuations can have different origins. For polar molecules with permanent dipoles, their orientational fluctuations lead to Keesom interaction. When orientational fluctuations of a permanent dipole induce a dipole in another nonpolar but polarizable molecule, the induced dipole leads to Debye interaction. In all remaining cases, transient dipoles of nonpolar polarizable molecules induce other transient dipoles and lead to London dispersion interactions.

In general, one can write the total vdW interaction potential  $V(\mathbf{r})$  between two molecules at positions  $\mathbf{r}_1$  and  $\mathbf{r}_2$  separated by  $\mathbf{r}=\mathbf{r}_1-\mathbf{r}_2$ , in the form

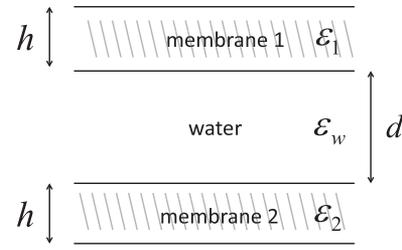
$$V(\mathbf{r}) = -\frac{\mathcal{C}}{r^6} = -\frac{3k_B T \alpha_1 \alpha_2}{(4\pi\epsilon_0)^2} \frac{1}{r^6}. \quad (6.127)$$

The  $1/r^6$  interaction reflects the dipolar nature of vdW interaction: dipolar field of the first molecule decays as  $1/r^3$ , interacts with the second molecule, and then propagates back to the first molecule, yielding a squared dipolar interaction,  $1/r^6$ . This argument does not take into account the relativistic corrections and, thus, corresponds to the nonretarded case.

The prefactor  $\mathcal{C}$  gives the strength of the interaction and is proportional to the product of polarizabilities,  $\alpha_1\alpha_2$ , of the two molecules and, thus, also to the product of molecular volumes. If the molecules interact in a medium of dielectric constant  $\epsilon_w$ , then the strength of the interaction is proportional to  $(\alpha_1\alpha_2)/\epsilon_w^2$ .

### 6.9.1 THE HAMAKER PAIRWISE SUMMATION

The simplest way to account for the vdW interactions between two large macroscopic bodies is called the *Hamaker summation*. It results from the pairwise summation of the



**FIGURE 6.17** Illustration of two planar membranes of thickness  $h$  each and with dielectric constants  $\epsilon_1$  and  $\epsilon_2$  interacting across a water slab of thickness  $d$  of dielectric constant  $\epsilon_w$ .

molecular vdW interactions, yielding for the interaction free energy:

$$\mathcal{F} = -\frac{3k_B T}{(4\pi\epsilon_0)^2} \int_{V_1} d^3 r_1 \int_{V_2} d^3 r_2 \frac{\alpha_1 n_1(\mathbf{r}_1) \alpha_2 n_2(\mathbf{r}_2)}{|\mathbf{r}_1 - \mathbf{r}_2|^6}, \quad (6.128)$$

where  $n_{1,2}$  are the densities of the molecules in the two bodies, and the volume integrals go over the volumes of the two bodies,  $V_1$  and  $V_2$ .

For two planar membranes of constant molecular density  $n_1$  and  $n_2$ , each of finite thickness  $h$  at a separation  $d$  (see Figure 6.17), the above Hamaker integral yields an interaction energy per unit surface area  $A$  of the form

$$\begin{aligned} \frac{\mathcal{F}(d, h)}{A} &= -\frac{3k_B T \alpha_1 n_1 \alpha_2 n_2}{(4\pi\epsilon_0 \epsilon_w)^2} \\ &\times \int_{d/2}^{d/2+h} dz_1 \int_{-d/2-h}^{-d/2} dz_2 \int_0^\infty \frac{2\pi\rho d\rho}{[(z_1 - z_2)^2 + \rho^2]^3}, \end{aligned} \quad (6.129)$$

where  $\mathbf{r}=(z, \rho)$  in cylindrical coordinates. In this case of two planar dielectric media (membranes) interacting across a gap of dielectric constant  $\epsilon_w$  (water), the excess polarizabilities are given by  $n_{1,2}\alpha_{1,2} = 2\epsilon_0\epsilon_w(\epsilon_{1,2} - \epsilon_w)/(\epsilon_{1,2} + \epsilon_w)$  (Israelachvili 2011). The integrals over  $z_1$  and  $z_2$  can be evaluated analytically, yielding

$$\frac{\mathcal{F}(d, h)}{A} = -\frac{\mathcal{H}}{12\pi} \left( \frac{1}{d^2} - \frac{2}{(d+h)^2} + \frac{1}{(d+2h)^2} \right), \quad (6.130)$$

with a prefactor defined as the *Hamaker constant*

$$\mathcal{H} = \frac{3k_B T}{4} \left( \frac{\epsilon_1 - \epsilon_w}{\epsilon_1 + \epsilon_w} \right) \left( \frac{\epsilon_2 - \epsilon_w}{\epsilon_2 + \epsilon_w} \right). \quad (6.131)$$

From this general expression, one can derive two interesting scaling limits for small and large inter-membrane separation,  $d$ .

For small separations,  $d \ll h$ , corresponding to the vdW interaction between two semi-infinite media separated by distance  $d$ ,

$$\frac{\mathcal{F}(d)}{A} \simeq -\frac{\mathcal{H}}{12\pi d^2} \sim \frac{1}{d^2}, \quad (6.132)$$

while for large separation,  $d \gg h$ , corresponding to the interaction of two thin sheets,

$$\frac{\mathcal{F}(d,h)}{A} \approx -\frac{\mathcal{H}h^2}{2\pi d^4} \sim \frac{1}{d^4}. \quad (6.133)$$

The  $d$  dependence obtained from the Hamaker summation is, to the lowest order, the same as obtained in more sophisticated approaches. However, the Hamaker constant,  $\mathcal{H}$ , can only be taken heuristically and, in fact, cannot be obtained from the simple pair-wise summation procedure.

## 6.9.2 MACROSCOPIC THEORY OF vdW INTERACTIONS

A more sophisticated theory for vdW interactions between macroscopic bodies was developed by J. M. Lifshitz in the 1950s (see Parsegian 2005). In the Lifshitz theory, the vdW interactions are *electromagnetic fluctuation interactions*, and the Hamaker coefficient,  $\mathcal{H}$ , is a functional of the frequency-dependent dielectric permeabilities of the interacting media. It can be evaluated from either experimentally determined dispersion properties or calculated dispersion spectra of the interacting materials. It consistently includes the relativistic retardation effects due to the finite velocity of light propagation and finite temperature effects (Woods et al. 2016).

In the Lifshitz theory, the free energy of two planar semi-infinite bodies is typically cast into the Hamaker-type form (Safran 1994, Parsegian 2005)

$$\frac{\mathcal{F}(d,h)}{A} = -\frac{\mathcal{H}(d,h)}{12\pi d^2}, \quad (6.134)$$

**TABLE 6.3**  
Values of the Hamaker Constant for Different Materials Interacting across Water (from Parsegian 2005)

Material	Hamaker Constant $\times 10^{21}$ [J]
Polystyrene	13
Polycarbonate	3.5
Hydrocarbons	3.8
Polymethyl methacrylate	1.47
Proteins	5–9

as in Eq. (6.132), but with the important difference that the  $d$ -dependent  $\mathcal{H}(d)$  can now be calculated explicitly via the Lifshitz formalism, when dielectric frequency-dependent properties of the interacting materials are available.

The most important characteristics of vdW interactions in the context of biomatter and, specifically, intermembrane interactions come from the presence of solvent, *i.e.*, water (Ninham and Parsegian 1970) and can be consistently taken into account within the Lifshitz theory. The calculated nonretarded Hamaker constant between lipid bilayers in water is found to be in the range of  $10^{-20} - 10^{-21}$  J. The high static dielectric constant of water and the low static dielectric constant of hydrocarbons (consisting of the membrane core region) lead to an anomalously large contribution to the entropic part of the vdW free-energy, which remains unretarded at all separations. Some characteristic values of the Hamaker constant for interaction of different materials across a water layer are given in Table 6.3.

## LIFSHITZ THEORY

The vdW interaction free energy is obtained in the Lifshitz theory as a sum, Eq. (6.136), over discrete imaginary Matsubara frequencies:

$$\zeta_n = \frac{2\pi k_B T n}{\hbar}, \quad n = 0, 1, 2, \dots, \quad (6.135)$$

where  $\hbar = h / 2\pi$  and  $h$  is the Planck constant. The terms of this sum involve the dielectric response functions of the different media: the two lipid bilayers separated by a slab of aqueous medium. The Hamaker constant in Eq. (6.134) depends on the dielectric response functions and can be obtained quantitatively.

The dielectric response function at imaginary frequencies,  $\varepsilon(i\zeta)$ , is given formally by the Kramers–Kronig relation (Smith 1985)

$$\varepsilon(i\zeta) = 1 + \frac{2}{\pi} \int_0^\infty \frac{\omega \varepsilon''(\omega)}{\omega^2 + \zeta^2} d\omega, \quad (6.136)$$

with  $\varepsilon''(\omega)$  being the imaginary part of the complex frequency-dependent dielectric function,  $\varepsilon(\omega) = \varepsilon'(\omega) + i\varepsilon''(\omega)$ . Quite generally,  $\varepsilon(i\zeta)$  is a real, monotonically decreasing function of its argument,  $\zeta$ . The Kramers–Kronig relation also establishes the connection between the vdW interactions and the measurable dispersion part of the dielectric response functions,  $\varepsilon''(\omega)$ . For this reason, the vdW interactions are also referred to as the *dispersion interactions*. The imaginary frequencies can be rationalized intuitively as follows: just as  $\varepsilon(\omega)$  characterizes the temporal response of a material to an external oscillating electric field  $\sim \exp(i\omega t)$ ,  $\varepsilon(i\zeta)$  characterizes the spontaneous time decaying fluctuation  $\sim \exp(-\zeta t)$ .

From the full Lifshitz formula for interacting lipid membranes, the limit of thick membranes,  $h \gg d$ , Eq. (6.134), without any retardation effects, yields the Hamaker constant in the form

$$\mathcal{H} = \frac{3}{2} k_B T \sum_{n=0}^{\infty} \prime [\Delta(i\zeta_n)]^2, \quad (6.137)$$

where the dielectric contrast is defined as

$$\Delta(i\zeta_n) = \frac{\varepsilon_L(i\zeta_n) - \varepsilon_w(i\zeta_n)}{\varepsilon_L(i\zeta_n) + \varepsilon_w(i\zeta_n)}, \quad (6.138)$$

in terms of the dielectric response function between the interacting media, *i.e.*, the lipid  $\varepsilon_L$  and water  $\varepsilon_w$ , at imaginary frequencies. Note that the prime in the summation of Eq. (6.137),  $\Sigma'$ , means that we have taken the lowest  $n = 0$  term with weight 1/2.

Standard forms for these dielectric responses can be used (Mahanty and Ninham 1976, Dagastine, Prieve and White 2000), where the dielectric response of water is described by twelve different relaxation frequencies: one microwave relaxation frequency, five infrared relaxation frequencies, and six ultraviolet relaxation frequencies. Similarly, the hydrocarbon materials (lipid membrane) are modelled by four  $\zeta$  ultraviolet relaxation frequencies (for details, see Parsegian 2005). In this limit of thick membranes,  $h \gg d$ , the Hamaker constant is indeed a constant, independent of the separation  $d$  but becomes  $d$ -dependent,  $\mathcal{H} = \mathcal{H}(d)$ , if the finite velocity of light is taken into account.

These retardation effects change the scaling of the interaction free energy from the  $1/d^2$  into a  $1/d^3$ , at separations of  $d \approx 10 - 100$  nm, usually too large to be of practical importance for interacting lipid membranes. We further stress that the summation over the discrete frequencies set is something that cannot be derived from a simple Hamaker summation procedure of Section 6.9.1.

Using model expressions for the dielectric response of the hydrocarbon core of lipid bilayers and the aqueous medium, one ends up with the value of  $4.3 \times 10^{-21}$  J for the relevant Hamaker coefficient (Podgornik, French and Parsegian 2006). A comparable value has been obtained in experiments for small membrane spacings. For example, it was found that the values of the Hamaker constant for dimyristoyl phosphatidylcholine (DMPC) and for dipalmitoyl phosphatidylcholine (DPPC) membranes, which forms a multistack, are in the range of  $2.87 - 9.19 \times 10^{-21}$  J (Petrache et al. 1998).

For thin membranes,  $h \ll d$ , the Lifshitz formula is valid in the nonretarded limit. The Hamaker constant of Eq. (6.137) has the same scaling of  $1/d^4$  as in the Hamaker summation of Eq. (6.133). The same Hamaker constant is obtained also for membranes of finite thickness (see Eq. 6.130).

VdW interactions for symmetric bodies, *e.g.*, two identical membranes interacting across a finite gap, are always

attractive, just as electrostatic interactions between two symmetric bodies in the PB theory are always repulsive (Neu 1999). For asymmetric bodies, *e.g.*, material '1' interacting with material '2' across water 'w', the Hamaker constant is given by a generalization of Eq. (6.137):

$$\mathcal{H} = \frac{3}{2} k_B T \sum_{n=0}^{\infty} \prime \Delta_{1w}(i\zeta_n) \Delta_{w2}(i\zeta_n). \quad (6.139)$$

This form of the Hamaker constant allows us to propose approximate *combining relations* that allow to extract unknown Hamaker constants from known ones (Israelachvili 2011). For the case of two media of material '1' interacting over material '3', the combining relation assumes the simple form  $\mathcal{H}_{131} \approx \mathcal{H}_{11} + \mathcal{H}_{33} - 2\mathcal{H}_{13}$ , where  $\mathcal{H}_{13}$  is for media '1' and '3' interacting across vacuum.<sup>2</sup>

The Lifshitz form of the Hamaker constant for asymmetric bodies also suggests that the vdW interaction can change sign, becoming repulsive. Sometimes this repulsive vdW interaction is referred to as quantum levitation (Munday, Capasso and Parsegian 2009).

### 6.9.3 THE DERJAGUIN–LANDAU–VERWEY–OVERBEEK THEORY

In the Derjaguin–Landau–Verwey–Overbeek (DLVO) theory (Verwey and Overbeek 1948), the total interaction energy between charged bodies (colloidal particles or membranes) is assumed to be a simple sum of the electrostatic and vdW interactions. For interacting planar bilayer membranes, the total interaction free energy is

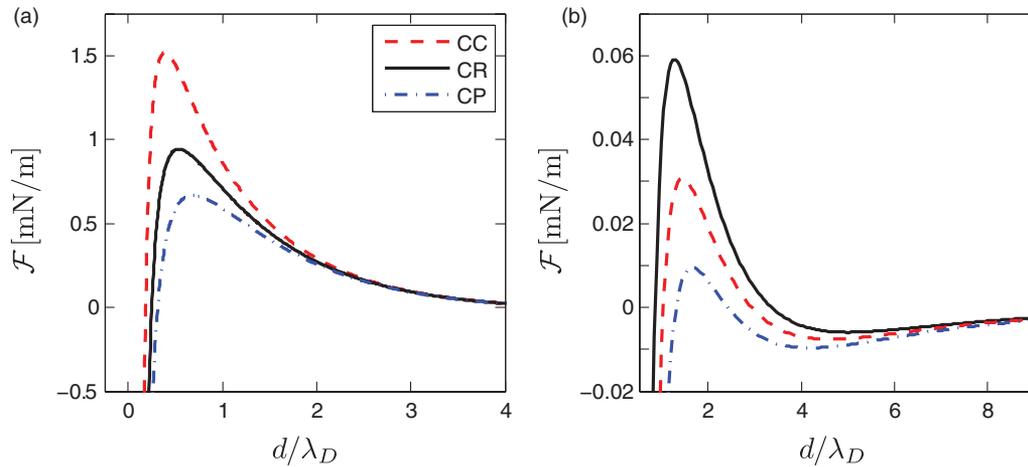
$$\mathcal{F}(d, h) = \mathcal{F}_{el}(d, h) + \mathcal{F}_{vdw}(d, h), \quad (6.140)$$

where the electrostatic part,  $\mathcal{F}_{el}(d, h)$ , is calculated in the PB framework as in Section 6.6. Since the PB osmotic pressure is easier to evaluate, the corresponding free-energy can be obtained via the integral

$$\frac{1}{A} \mathcal{F}_{el}(d, h) = \int_d^{\infty} d\ell \Pi(\ell), \quad (6.141)$$

with  $\Pi$  calculated as in Section 6.8. The vdW interaction free energy,  $\mathcal{F}_{vdw}$ , is calculated either from the Hamaker summation procedure or, more appropriately, from the Lifshitz theory.

Because of the  $1/d^2$  dependence of the vdW free energy, the total interaction free energy exhibits a universal *primary minimum* at vanishing spacings. However, since the assumption of the continuum solvent is bound to break down in this small  $d$  limit, this minimum is more related to the continuum theory assumption inherent in the Lifshitz theory than to the physical interactions between membranes. In fact, a strong *hydration interaction* (related to the breakdown of the continuum model of the solvent) usually obliterates the primary minimum and results in a



**FIGURE 6.18** The interaction free energy density for two symmetric membranes,  $\mathcal{F}(d, h) / A$ , Eq. (6.140). In (a), the electrostatic part is calculated with three different boundary conditions: constant charge (CC), constant potential (CP), and charge regulation (CR). The inequality,  $\mathcal{F}_{CC} > \mathcal{F}_{CR} > \mathcal{F}_{CP}$ , seen here is a general relation. The parameters used are  $a = 0.5$  nm,  $n_b = 0.1$  M, and  $\alpha = -6$  (pK  $\approx 1.48$ ). In (b), we calculate the electrostatic part of the free energy for constant charge boundary conditions with  $\sigma = -e/(7 \text{ nm}^2)$  for three different bulk ionic concentrations,  $n_b = 0.32$  M (solid black line),  $n_b = 0.35$  M (dotted red line), and  $n_b = 0.38$  M (dash-dotted blue line). The membrane width in (a) and (b) is  $h = 4$  nm.

monotonic repulsive interaction even for very small intermembrane separations (Parsegian, Fuller and Rand 1979).

Apart from the primary minimum at small separations, a secondary minimum can emerge at larger separations, depending on the system parameters. In Figure 6.18, we plot some of these scenarios for various assumptions on the electrostatic interactions. The curves shown in Figure 6.18 embody the essence of the DLVO theory. They can either show a monotonic repulsion extending over the whole ranges of separation, or attraction that is turned into repulsion or vice versa. In Figure 6.18a, the DLVO free energy is shown for the three different boundary conditions: constant charge (CC), constant potential (CP), and charge regulation (CR). Two minima are clearly seen: a minimum at  $d / \lambda_D \rightarrow 0$ , which is the nonphysical primary minimum mentioned earlier, and a minimum at  $d / \lambda_D \rightarrow \infty$ .

In Figure 6.18b, we show the DLVO free energy for CC boundary conditions with three different ionic concentrations:  $n_b = 0.32$  M ( $\lambda_D \approx 0.54$  nm),  $n_b = 0.35$  M ( $\lambda_D \approx 0.52$  nm), and  $n_b = 0.38$  M ( $\lambda_D \approx 0.5$  nm). As shown in the figure, for these bulk salt concentrations, a secondary, very shallow, minimum appears. Increasing  $n_b$  lowers the energy barrier and strengthens the shallow secondary minimum. The appearance of this secondary minimum is an essential ingredient in the explanation of the stability of colloidal particles and interacting membranes. They come into stable equilibrium at this secondary minimum. The high energy barrier between the secondary and the primary minimum makes it stable.

## 6.10 LIMITATIONS AND GENERALIZATIONS

The DLVO theory relies on approximations that have a finite range of validity (Naji et al. 2013). First, the vdW interaction is not really decoupled from the PB mean-field

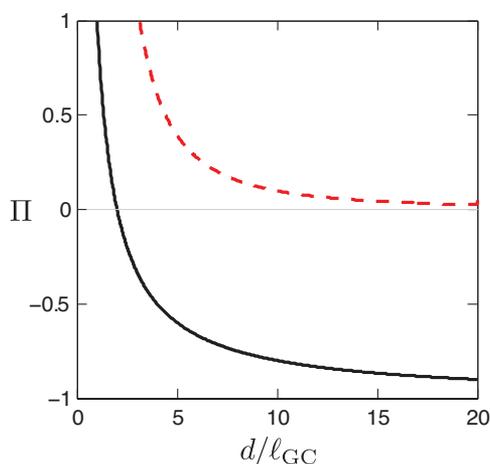
formulation and should be correspondingly modified. And second, one of the central results of PB theory that symmetric bodies always repel each other is *incorrect*. For physically interesting situations involving highly charged interfaces, or multivalent mobile ions, the electrostatic interaction can, in fact, be attractive.

These drawbacks of the classical DLVO theory, describing interactions between charged colloidal bodies or interacting membranes, can be amended. Recently, a new paradigm introduced a transparent systematization of the electrostatic interactions between charged bodies in terms of two useful regimes: *weak coupling* (WC) and *strong coupling* (SC) (Boroudjerdi et al. 2005). This allows a more accurate evaluation of the electrostatic interactions and the coupling between electrostatic and vdW interactions.

In order to introduce the WC and SC regimes, one needs to consider the relative strength of electrostatic interactions as compared with the background thermal energy. The thermal energy can be compared either with Coulomb interaction between two  $q = ze$  charges of valency  $z$  giving rise to a modified Bjerrum length,  $z^2 \ell_B$ , as well as a modified Gouy–Chapman length,  $\ell_{GC}/z$ , quantifying the strength of the electrostatic interaction between a point charge ( $q = ze$ ) and a surface charge density,  $\sigma$ . Dividing the two lengths leads to a fundamental dimensionless *electrostatic coupling parameter* introduced by Netz (2001):

$$\Xi = z^2 \ell_B / (\ell_{GC} / z) = 2\pi z^3 \ell_B^2 |\sigma|. \quad (6.142)$$

For a system composed purely of counterions, the regimes of WC and SC can be understood in the following way. When the coupling parameter is small, *i.e.*,  $\Xi \ll 1$ , one goes back to the PB theory, with an addition of thermal fluctuations contributing a vdW-like interaction that scales linearly with  $\Xi$  and partially replaces the Lifshitz theory results.



**FIGURE 6.19** The pressure for two symmetric membranes in units of  $k_B T / (2\pi\ell_B \ell_{GC}^2)$  as function of the dimensionless distance  $d/\ell_{GC}$ . The calculation is done for the counterions-only case as obtained from the PB theory (dashed red line), Eqs. (6.65) and (6.66), in comparison with the strong coupling (SC) result (solid black line),  $\Pi = -1 + 2\ell_{GC}/d$  in units of  $k_B T / (2\pi\ell_B \ell_{GC}^2)$  (Ghodrat et al. 2015).

This clearly establishes a connection between electrostatic and vdW interactions.

In the opposite limit of large  $\Xi \gg 1$ , one observes a very different behaviour and important deviations from the standard DLVO paradigm. Here, the mobile ions become strongly correlated. It leads to a fundamental consequence that the interactions between nominally equally charged surfaces can become *attractive*. This is shown clearly in Figure 6.19 that compares the WC and SC results for the osmotic pressure between two equally charged membranes. Extensive Monte-Carlo simulations show that the WC to SC crossover is associated with a jump in the heat capacity and the appearance of short-range correlations between counterions for the coupling range  $10 < \Xi < 100$  (Naji et al. 2013). For very high values of the coupling parameter,  $\Xi \approx 3 \times 10^4$ , a transition to a Wigner crystalline phase (characterized by a diverging heat capacity) occurs. Therefore, the whole DLVO idea that the interaction is composed of repulsive electrostatic interaction and an attractive vdW, needs a serious revision for  $\Xi \gg 1$  (SC regime).

The WC/SC paradigm is clear for the simple counterions-only case but becomes more complex in the presence of additional mobile charge components, *e.g.*, polyvalent plus monovalent salt, surface charge heterogeneity, mobile ion with multipolar structure, and polarizable mobile ions. These different components introduce new coupling parameters (similar to  $\Xi$ ) that lead to additional features making the general form of the electrostatic interactions more complicated than in the simple PB and DLVO framework (Podgornik and Andelman 2020).

In this review, we have presented in detail the PB treatment for mobile ions in solutions for planar geometries as applicable to charged membranes. The PB equation is a mean-field equation that is obtained from the zeroth

approximation in the WC regime (Podgornik and Žekš 1988, Borukhov, Andelman and Orland 1998, 2000, Netz and Orland 2000, Markovich, Andelman and Podgornik 2014, 2015).

Our chapter does not treat the fluctuations around the mean-field solution explicitly but only indirectly via the vdW interactions. Thermal fluctuations, which represent ion–ion correlations, have a key role in many interesting phenomena. Among them, we mention surface tension of electrolyte solutions (Onsager and Samaras 1934, Markovich, Andelman and Podgornik 2014, 2015, 2016, 2017a) and their dielectric decrement (Ben-Yaakov, Andelman and Podgornik 2011, Levy, Andelman and Orland 2012, Adar, Markovich and Andelman 2017, Adar et al. 2018). Furthermore, we do not cover the SC regime but only give some of its interesting results in this last section.

Apart from a more accurate treatment of electrostatic interactions, other simplifying assumptions were made at the base of the PB and DLVO theories presented in this chapter. As noted in the beginning of the chapter, the membrane structure was completely ignored. Therefore, membrane heterogeneities, curvature, and undulations were not discussed. The solvent (water) was treated as a featureless media with dielectric constant  $\epsilon_w$ , and the effect of the mobile ions and the solvent structure (*e.g.*, water permanent dipole) on the decrement of the solvent dielectric constant was ignored (Ben-Yaakov, Andelman and Podgornik 2011, Levy, Andelman and Orland 2012, Adar et al. 2018). This effect also gives rise to a dielectrophoretic saturation of the counterions close to the membrane, similar to the steric mPB saturation in Section 6.4 (Nakayama and Andelman 2015).

Other extensions of the PB theory can be done by considering mixture of solvents (Ben-Yaakov et al. 2009a and 2009b) or nonelectrostatic interaction between the mobile ions themselves such as hydration interactions (Burak and Andelman 2000) or between the membrane and the mobile ions (Markovich, Andelman and Podgornik 2014, 2015, 2016, 2017a). For simplicity, the ions were treated as point-like particles which neglects their internal structure and polarizability (see, *e.g.*, Démery, Dean and Podgornik 2012 and references therein). Unfortunately, these interesting developments lie beyond the scope of this chapter and will be covered elsewhere (Andelman, Burak, and Orland, to be published).

## NOTES

- 1 Throughout this chapter we use the SI unit system.
- 2 Note that the *combining relations* can be also obtained from the simpler Hamakar pair-wise summation of Section 6.9.1.

## BIBLIOGRAPHY

- Adamson, A. W. and A. P. Gast. *Physical Chemistry of Surfaces*, 6th ed. New York: Wiley, (1997).
- Adar, R. M., T. Markovich and D. Andelman. “Bjerrum pairs in ionic solutions: A Poisson–Boltzmann approach”. *J. Chem. Phys.* (2017): 146, 194904.

- Adar, R. M., T. Markovich, A. Levy, H. Orland and D. Andelman. "Bjerrum pairs in ionic solutions: A Poisson-Boltzmann approach". *J. Chem. Phys.* (2018): 149, 054504.
- Andelman, D., "Electrostatic properties of membranes: The Poisson-Boltzmann theory". In *Handbook of Physics of Biological Systems*, edited by Lipowsky R. and Sackman E., Vol. I. Amsterdam: Elsevier Science, (1995), Chap. 12.
- Andelman, D., "Introduction to electrostatics in soft and biological matter", In *Soft Condensed Matter Physics in Molecular and Cell Biology*, Ed. by Poon W. and Andelman D., Scottish Graduate Series: SUSSP 59. New York: Taylor and Francis, (2005), pp. 97–122.
- Andelman, D., Y. Burak and H. Orland. *Electrostatic Interactions in Soft and Biological Matter: Ions, Membranes, Polymers and Colloids*. New York: Cambridge University Press, to be published.
- Ariel, G., H. Diamant and D. Andelman. "Kinetics of surfactant adsorption at fluid-fluid interfaces: Surfactant mixtures". *Langmuir* (1999): 15, 3574–3581.
- Avni, Y., D. Andelman and R. Podgornik. "Charge regulation with fixed and mobile charged macromolecules". *Curr. Opin. Electrochem.* (2019): 13, 70–77.
- Avni, Y., T. Markovich, R. Podgornik and D. Andelman. "Charge regulating macro-ions in salt solutions: Screening properties and electrostatic interactions". *Soft Matter*. (2018): 14, 6058–6069.
- Avni, Y., R. Podgornik and D. Andelman. "Critical Behavior of Charged-regulated Macro-ion". *J. Chem. Phys.* (2020): 153, 024901
- Ben-Yaakov, D. and D. Andelman. "Revisiting the Poisson-Boltzmann theory: Charge surfaces, multivalent ions and inter-plate forces". *Physica A* (2010): 389, 2956–2961.
- Ben-Yaakov, D., D. Andelman, D. Harries and R. Podgornik. "Ions in mixed dielectric solvents: Density profiles and osmotic pressure between charged interfaces". *J. Phys. Chem. B* (2009a): 113, 6001–6011.
- Ben-Yaakov, D., D. Andelman, D. Harries and R. Podgornik. "Beyond standard Poisson-Boltzmann Theory: Ion-specific interactions in aqueous solutions". *J. Phys.: Condens. Matter* (2009b): 21, 424106.
- Ben-Yaakov, D., D. Andelman and R. Podgornik. "Dielectric decrement as a source of ion-specific effects". *J. Chem. Phys.* (2011): 134, 074705.
- Ben-Yaakov, D., Y. Burak, D. Andelman and S. A. Safran. "Electrostatic interactions of asymmetrically charged membranes". *Europhys. Lett.* (2007): 79, 48002.
- Bordag, M., G. L. Klimchitskaya, U. Mohideen, and V. M. Mostapenko. *Advances in the Casimir Effect*. Oxford: Oxford University Press, 2009.
- Borkovec, M., B. Jönsson and G. J. M. Koper. "Ionization processes and proton binding in polyprotic systems: Small molecules, proteins, interfaces and polyelectrolytes". *Surf. Colloid Sci.* (2001): 16, 99–339.
- Boroudjerdi, H., Y. W. Kim, A. Naji, R. R. Netz, X. Schlagberger and A. Serr. "Statics and dynamics of strongly charged soft matter". *Phys. Rep.* (2005): 416 129–199.
- Borukhov, I., D. Andelman and H. Orland. "Steric effects in electrolytes: A modified Poisson-Boltzmann equation". *Phys. Rev. Lett.* (1997): 79, 435–438.
- Borukhov, I., D. Andelman and H. Orland. "Random polyelectrolytes and polyampholytes in solution". *Eur. Phys. J. B* (1998): 5, 869–880.
- Borukhov, I., D. Andelman and H. Orland. "Adsorption of large ions from an electrolyte solution: A modified Poisson-Boltzmann equation". *Electrochimica Acta* (2000): 46, 221–229.
- Burak, Y. and D. Andelman. "Hydration interactions: Aqueous solvent effects in electric double layers". *Phys. Rev. E* (2000): 62, 5296–5312.
- Chapman, D. L. "A contribution to the theory of electrocapillarity". *Philos. Mag* (1913): 25, 475.
- Churaev, N. V., B. V. Derjaguin and V. M. Muller. *Surface Forces*. New York: Springer, 2014.
- Dagastine R. R., D. C. Prieve and L. R. White. "The dielectric function for water and its application to van der Waals forces". *J. Colloid Interface Sci.* (2000): 231, 351–358.
- Davies, J. T. "Adsorption of long-chain ions. I". *Proc. R. Soc. A* (1958): 245, 417–428.
- Dean D. S., J. Dobnikar, A. Naji, and R. Podgornik, eds. *Electrostatics of Soft and Disorder Matter*. CRC Press, 2014.
- Dean, D. S. and R. R. Horgan. "Field theoretic derivation of the contact value theorem in planar geometries and its modification by the Casimir effect". *Phys. Rev. E* (2003): 68, 061106.
- Debye, P. and E. Hückel. "Zur Theorie der Elektrolyte. Gefrierpunktniedrigung und verwandte Erscheinungen. (The theory of electrolytes. Lowering of freezing point and related phenomena)". *Physik Z.* (1923): 24, 185–206.
- Démery, V., D. S. Dean and R. Podgornik. "Electrostatic interactions mediated by polarizable counterions: Weak and strong coupling limits". *J. Chem. Phys.* (2012): 137, 174903.
- Diamant, H. and D. Andelman. "Kinetics of surfactant adsorption at fluid-fluid interfaces". *J. Phys. Chem.* (1996): 100, 13732–13742.
- Evans, D. F. and H. Wennerström, 1999. *The Colloidal Domain*, 2nd edition. New York: VCH Publishers, 1999.
- Ghodrat, M., A. Naji, H. Komaie-Moghaddam and R. Podgornik. "Strong coupling electrostatics for randomly charged surfaces: Antifragility and effective interactions". *Soft Matter*. (2015): 11, 3441–3459.
- Gouy, G. "Sur la constitution de la charge électrique à la surface d'un électrolyte". *J. Phys. (France)* (1910): 9, 457–468.
- Gouy, G. "Sur la fonction électrocapillaire". *Ann. Phys. (Leipzig)* (1917): 7, 129–184.
- Grahame, D. C. "The electrical double layer and the theory of electrocapillarity". *Chem. Rev.* (1947): 41, 441–501.
- Hallett, J. E., D. A. Gillespie, R. M. Richardson and P. Bartlett. "Charge regulation of nonpolar colloids". *Soft Matter* (2018): 14, 331–343.
- Henderson, D. and L. Blum. "Some comments regarding the pressure tensor and contact theorem in a nonhomogeneous electrolyte". *J. Chem. Phys.* (1981): 75, 2025–2026.
- Holm, C., P. Kékicheff and R. Podgornik, eds. *Electrostatic Effects in Soft Matter and Biophysics*. Amsterdam: Kluwer Academic Press, 2000.
- Israelachvili, J. N. *Intermolecular and Surface Forces, 3rd edition*. London: Academic Press, 2011.
- Kanduč, M., M. Trulsson, A. Naji, Y. Burak, J. Forsman and R. Podgornik. "Weak- and strong-coupling electrostatic interactions between asymmetrically charged planar surfaces". *Phys. Rev. E* (2008): 78, 061105.
- Kilic, M. S., M. Z. Bazant and A. Ajdari. "Steric effects in the dynamics of electrolytes at large applied voltages: I. Double-layer charging". *Phys. Rev. E* (2007): 75, 021502.
- Kiometzis, M. and H. Kleinert. "Electrostatic stiffness properties of charged bilayers". *Phys. Lett. A* (1989): 140, 520–524.
- Kornyshev, A. A. "Double layer in ionic liquids: Paradigm change?" *J. Phys. Chem. B* (2007): 111, 5545–5557.
- Lau, A. W. C. and P. Pincus. "Binding of oppositely charged membranes and membrane reorganization". *Eur. Phys. J. B* (1999): 10, 175–180.

- Levy, A., D. Andelman and H. Orland. "Dielectric constant of ionic solutions: A field-theory approach". *Phys. Rev. Lett.* (2012): 108, 227801.
- Mahanty, J. and B. W. Ninham. *Dispersion Forces*. London: Academic Press, 1976.
- Markovich, T., D. Andelman and H. Orland. "Ionic profiles close to dielectric discontinuities: Specific ion-surface interactions". *J. Chem. Phys.* (2016): 145, 134704.
- Markovich, T., D. Andelman and R. Podgornik. "Surface tension of electrolyte solutions: A self-consistent theory". *Europhys. Lett.* (2014): 106, 16002.
- Markovich, T., D. Andelman and R. Podgornik. "Surface tension of electrolyte interfaces: Ionic specificity within a field-theory approach". *J. Chem. Phys.* (2015): 142, 044702.
- Markovich, T., D. Andelman and R. Podgornik. "Charge regulation: A generalized boundary condition?" *EPL* (2016): 113, 26004.
- Markovich, T., D. Andelman and R. Podgornik. "Surface tension of acid solutions: Fluctuations beyond the nonlinear Poisson-Boltzmann theory". *Langmuir* (2017a): 33, 34–44.
- Markovich, T., D. Andelman and R. Podgornik. "Complex fluids with mobile charge-regulating macro-ions". *EPL* (2017b): 120, 26001.
- Munday, J. N., F. Capasso and V. A. Parsegian. "Measured long-range repulsive Casimir-lifshitz forces." *Nature* (2009): 457, 170–173.
- Naji, A., M. Kanduč, J. Forsman and R. Podgornik. "Perspective: Coulomb fluids - weak coupling, strong coupling, in between and beyond". *J. Chem. Phys.* (2013): 139, 150901.
- Nakayama, Y. and D. Andelman. "Differential capacitance of the electric double layer: The interplay between ion finite size and dielectric decrement". *J. Chem. Phys.* (2015): 142, 044706.
- Netz, R. R. "Electrostatistics of counter-ions at and between planar charged walls: From Poisson-Boltzmann to the strong-coupling theory". *Eur. Phys. J. E* (2001): 5, 557–574.
- Netz, R. R. and H. Orland. "Beyond Poisson-Boltzmann: Fluctuation effects and correlation functions". *Eur. Phys. J. E* (2000): 1, 203–214.
- Neu, J. C. "Wall-mediated forces between like-charged bodies in an electrolyte". *Phys. Rev. Lett.* (1999): 82, 1072–1074.
- Ninham, B. W. and V. A. Parsegian. "Van der Waals forces: Special characteristics in lipid-water systems and a general method of calculation based on the lifshitz theory". *Biophys. J.* (1970): 10, 646–663.
- Ninham, B. W. and V. A. Parsegian. "Electrostatic potential between surfaces bearing ionizable groups in ionic equilibrium with physiologic saline solution". *J. Theor. Biol.* (1971): 31, 405–428.
- Onsager, L. and N. N. T. Samaras. "The surface tension of Debye-Hückle electrolytes". *J. Chem. Phys.* (1934): 2, 528–536.
- Parsegian V. A. *Van der Waals Forces*. New York: Cambridge University Press, 2005.
- Parsegian V. A., N. Fuller and R. P. Rand. "Measured work of deformation and repulsion of lecithin bilayers". *Proc. Natl. Acad. Sci. USA* (1979): 76, 2750–2754.
- Parsegian V. A. and D. Gingell. "On the electrostatic interaction across a salt solution between two bodies bearing unequal charges". *Biophys. J.* (1972): 12, 1192–1204.
- Petrache, H. I., N. Gouliarov, S. Tristram-Nagle, R. Zhang, R. M. Suter, and J. F. Nagle. "Interbilayer interactions from high-resolution X-Ray scattering". *Phys. Rev. E* (1998): 57, 7014–7024.
- Podgornik, R. and D. Andelman. "Embarras de Richesses in non-DLVO Colloidal Interactions". *J. Club Cond. Matt. Phys.* <https://arxiv.org/abs/2101.00187> (2020).
- Podgornik, R., R. H. French and V. A. Parsegian. "Nonadditivity in van der Waals interactions within multilayers". *J. Chem. Phys.* (2006): 124, 044709.
- Podgornik, R. and B. Zeks. "Inhomogeneous coulomb fluid: A functional integral approach". *J. Chem. Soc., Faraday Trans. 2*(1988): 84, 611–631.
- Sader, J. E. and D. Y. C. Chan. "Electrical double layer interaction between charged particles near surfaces and in confined geometries". *J. Colloid Interface Sci.* (1999): 213, 268–269.
- Safran, S. A. *Statistical Thermodynamics of Surfaces, Interfaces, and Membranes*. Boston, MA: Addison-Wesley, 1994.
- Verwey, E. J. W. and J. T. G. Overbeek. *Theory of the Stability of Lyophobic Colloids*. New York: Elsevier, 1948.
- Winterhalter, M. and W. Helfrich. "Bending elasticity of electrically charged bilayers: Coupled monolayers, neutral surfaces, and balancing stresses". *J. Phys. Chem.* (1992): 96, 327–330.
- Woods, L. M., D. A. R. Dalvit, A. Tkatchenko, P. Rodriguez-Lopez, A. W. Rodriguez and R. Podgornik. "Materials perspective on Casimir and van der Waals interactions," *Rev. Mod. Phys.* (2016): 88, 045003.

**Corrections to “Charged Membranes” (Chapter 6) in:  
“Handbook of Lipid Membranes”**

by T. Markovich, D. Andelman and R. Podgornik

1. Page 108, Eq. (6.41) should read:

$$S_v = k_B \ln \left( \frac{N!}{N_+! N_-! N_w!} \right) = -k_B [ \dots ] \quad (6.41)$$

2. Page 108, Eq. (6.42) should read:

$$S_V = \int \frac{d^3 r}{v} S_v = -k_B \int d^3 r [ \dots ] \quad (6.42)$$

3. Page 111, Eq. (6.55) should read:

$$\begin{aligned} \Pi &= -\frac{k_B T}{8\pi l_B} (\Psi')^2 + k_B T \sum_{i=1}^M (n_i(z) - n_i^{(b)}) \\ &= \text{const.} \end{aligned} \quad (6.55)$$